

AD-A104 836

CALIFORNIA UNIV BERKELEY DEPT OF CHEMISTRY
INTERACTION POTENTIALS BETWEEN RARE GAS ATOMS AND HALOGEN ATOMS--ETC(U)
JUN 81 Y T LEE

F/G 20/7
N00014-77-C-0101

NL

UNCLASSIFIED

1
2
3
4
5
6
7
8
9
10
11
12
13
14
15
16
17
18
19
20
21
22
23
24
25
26
27
28
29
30
31
32
33
34
35
36
37
38
39
40
41
42
43
44
45
46
47
48
49
50
51
52
53
54
55
56
57
58
59
60
61
62
63
64
65
66
67
68
69
70
71
72
73
74
75
76
77
78
79
80
81
82
83
84
85
86
87
88
89
90
91
92
93
94
95
96
97
98
99
100

END
FILE
FILED
JUN 81
DSC

AD A104836

11BS

(12)

SECURITY CLASSIFICATION OF THIS PAGE (When Data Entered)

REPORT DOCUMENTATION PAGE		READ INSTRUCTIONS BEFORE COMPLETING FORM
1. REPORT NUMBER NR0395-572	2. GOVT ACCESSION NO.	3. RECIPIENT'S CATALOG NUMBER
4. TITLE (and Subtitle) INTERACTION POTENTIALS BETWEEN RARE GAS ATOMS AND HALOGEN ATOMS		5. TYPE OF REPORT & PERIOD COVERED Final Technical Report 12/1/77-12/1/80
7. AUTHOR(s) Professor Yuan T. Lee		6. PERFORMING ORG. REPORT NUMBER N00014-77-C-0101
8. CONTRACT OR GRANT NUMBER(s)		10. PROGRAM ELEMENT, PROJECT, TASK AREA & WORK UNIT NUMBERS
9. CONTROLLING OFFICE NAME AND ADDRESS Dr. W.J. Condell, Director, Physics Program, Department of the Navy Office of Naval Research, Arlington, VA		12. REPORT DATE June 10, 1981
11. MONITORING AGENCY NAME & ADDRESS (if different from Controlling Office)		13. NUMBER OF PAGES 73
		15. SECURITY CLASS. (of this report) Unclassified
		15a. DECLASSIFICATION/DOWNGRADING SCHEDULE
16. DISTRIBUTION STATEMENT (of this Report) DISTRIBUTION UNLIMITED SEP 30 1981		
17. DISTRIBUTION STATEMENT (of the document entered in Block 20, if different from Report) DISTRIBUTION STATEMENT A Approved for public release; Distribution Unlimited		
18. SUPPLEMENTARY NOTES		
19. KEY WORDS (Continue on reverse side if necessary and identify by block number) Interaction potentials between rare gas atoms and halogen atoms. Crossed molecular beams measurements of elastic scattering. Cross sections for rare gas-halogen atom systems.		
20. ABSTRACT (Continue on reverse side if necessary and identify by block number) Interaction potentials are investigated for F + Ne, Ar, Kr, Xe, Cl + Xe and Br + Ar, Kr, Xe from the measurements of elastic differential cross sections in crossed molecular beams experiments. The general trend of the transition from van der Waals interactions to chemical interactions is carefully examined.		

FILE COPY

DD FORM 1 JAN 73 1473

EDITION OF 1 NOV 68 IS OBSOLETE

S/N 0102-LF-014-6601

400375

SECURITY CLASSIFICATION OF THIS PAGE (When Data Entered)

817 30050

FINAL TECHNICAL REPORT PREPARED FOR ONR
FOR CONTRACT N00014-77-C-0101

TITLE: INTERACTION POTENTIALS BETWEEN RARE
GAS ATOMS AND HALOGEN ATOMS

PRINCIPAL INVESTIGATOR

Professor Yuan T. Lee
Department of Chemistry
University of California
Berkeley, CA 94720

PERIOD COVERED

December 1, 1977 to December 1, 1980

INTRODUCTION

This contract deals with the research on the interaction potentials between rare gas atoms with halogen molecules by the crossed molecular beams method. Detailed information on ground state interactions between F, Cl, Br and rare gas atoms are derived from the measurements of elastic differential scattering cross sections.

In the first two years of the contract period, the systems including F and Cl were investigated. During the final year, the systems of Br + Ar, Kr and Xe were studied.

Because of the drastic cut of the budget for the third year, from the planned \$55,000 to \$30,000, the project has not been brought to conclusion in an orderly manner. Funds from other sources had to be sought to complete the works and which delayed the completion of this research as well as the completion of this final report. The failure of honoring the original budget which had been agreed upon at the initiation of the project is very disruptive to the healthy operation of a research laboratory, it should be avoided at any cost.

In this final report, the work on Br(2P) + Ar, Kr, Xe(1S) is included.

Accession For	NTIS GR-41	DTIC TAB	Announced	Justification for Release	By _____	Distribution	Availability Cross	Available and/or	Special
---------------	------------	----------	-----------	---------------------------	----------	--------------	--------------------	------------------	---------

PUBLICATIONS

1. Crossed Molecular Beam Studies on the Interaction Potential for $F(^2P) + Xe(^1S)$. C. H. Becker, P. Casavecchia and Y. T. Lee, J. Chem. Phys. 69, 2377 (1978).

Abstract

For the evaluation of the ground state XeF interaction potential, angular distributions of F atoms scattered off Xe were measured in crossed molecular beam experiments at collision energies of 2.11, 10.5, and 13.9 kcal/mole. F atoms, produced by thermal dissociation of F_2 at 700°C in a supersonic expansion using rare gas carriers, contain ~78% F ($^2P_{3/2}$) ground state and ~22% F ($^2P_{1/2}$) spin-orbit excited state. Consequently, three electronic states, $X1/2$, $I3/2$ emerging from the $^2P_{3/2} + ^1S_0$ asymptote, and $III1/2$ from the $^2P_{1/2} + ^1S_0$ asymptote are involved in the scattering. A simple elastic approximation, neglecting interstate coupling, is used for the calculation of differential scattering cross sections in the evaluation of interaction potentials. Experimental results are found corroborating the spectroscopically derived potential of Tellinghuisen et al ($\epsilon = 3.359$ kcal/mole, $r_m = 2.293 \text{ \AA}$) for $V_{X1/2}(r)$, and the conclusion that $V_{I3/2}(r)$ and $V_{III1/2}(r)$ are in close resemblance to the ground state Ne + Xe interaction potential.

2. Crossed Molecular Beam Studies on the Interaction Potentials for $Cl(^2P) + Xe(^1S)$. C. H. Becker, J. J. Valentini, P. Casavecchia, S. J. Sibener and Y. T. Lee, Chem. Phys. Lett. 61, 1 (1979).

Abstract

A crossed beam experiment has provided angular distributions of $^{35}Cl(^2P_{3/2,1/2})$ scattered off $Xe(^1S_0)$ at 4 collision energies from 2.37 to 26.1 kcal/mole. Interaction potentials for the $X1/2$, $I3/2$, and $III1/2$ states are obtained via an approximate elastic scattering analysis as done previously for F-Xe, Kr, Ar, Ne. The $X1/2$ potential obtained ($\epsilon = 0.80$ kcal/mole, $r_m = 3.23 \text{ \AA}$) is a corroboration of the recent spectroscopically derived potential of Sur et al. The $I3/2$ potential ($\epsilon = 0.37$ kcal/mole, $r_m = 4.1 \text{ \AA}$) closely resembles that of ground state Ar-Xe.

3. Crossed Molecular Beam Studies on the Interaction Potentials for F(²P) + Ne, Ar, Kr(¹S). C. H. Becker, P. Casavecchia and Y. T. Lee, J. Chem. Phys. 70, 2986 (1979).

Abstract

Angular distributions are presented of F(²P_{3/2,1/2}) scattered off Ne, Ar, and Kr in the thermal energy range measured in crossed molecular beams experiments. The data are analyzed as previously done for the F-Xe system to obtain the interaction potential curves for the three relevant states (X 1/2, I 3/2, II 1/2) governing the scattering. The resulting X 1/2 curves for the fluorine-rare gas systems show no simple trend and the origin of these interactions is discussed. The I 3/2 and II 1/2 states resemble those of the ground state Ne-Ne, Ar, Kr systems.

4. Coupled-Channel Study of Halogen (²P) + Rare Gas (¹S) Scattering. C. H. Becker, P. Casavecchia, Y. T. Lee, R. E. Olson, and W. A. Lester, Jr., J. Chem. Phys. 70, 5477 (1979).

Abstract

Quantum mechanical coupled-channel (CC) scattering calculations are reported using realistic adiabatic potentials for the ²P + ¹S interaction of F-Ar, F-Xe, and Cl-Xe. Differential cross sections $d\sigma/d\omega$ derived from a simple elastic approximation appropriate for large spin orbit interactions accurately reproduce all the gross features computed by the coupled-channel method. This finding supports the extraction of interaction potentials from laboratory differential cross sections $I(\theta)$ via an elastic analysis. Integral inter and intramultiplet changing cross sections are expressed conveniently in terms of Grawert's $B(j, j'; g)$ coefficient. Information on the collision dynamics is extracted by following the partial wave dependence of selected $B(j, j'; g)$. Classical turning point analysis, based on the values of the large l-waves for which these partial wave contributions $B_l(j, j'; g)$ begin to rise above zero, leads to the conclusion that both intermultiplet and first order forbidden intramultiplet transitions are caused by a single localized nonadiabatic coupling region at the position of complex crossing of the $\Omega = 1/2$ adiabatic potentials. Small amplitude oscillations or perturbations in the CC calculated $d\sigma/d\omega$ and in the experimental $I(\theta)$ are thought to be examples of Stükelberg oscillations, though

quantitative agreement between these quantities is not obtained. The energy dependence and interference structure of the computed $B(j, j; g)$ are briefly discussed, as is the approximation of the constant spin orbit interaction over the experimentally accessible range of internuclear distances.

5. Interaction Potentials of Ground State Xenon Halides. P. Casavecchia, Guozhong He, R. K. Sparks, and Y. T. Lee, Proceedings of the 8th International Symposium on Molecular Beams, Palais des Festivals, Cannes, France, June 1-5, 1981, LBL-11373 ABS (attached).
6. Interaction Potentials for $\text{Br}({}^2\text{P}) + \text{Ar}, \text{Kr}, \text{Xe}({}^1\text{S})$ by the Crossed Beams Method. Piergiorgio Casavecchia, Guozhong He, Randal K. Sparks, and Yuan T. Lee, J. Chem. Phys. (1981) in press, LBL-11907. This article is attached together with this final report.

BUDGET STATEMENT

December 1, 1977 to November 30, 1978	\$ 50,000
December 1, 1978 to November 30, 1979	\$ 50,000
December 1, 1979 to November 30, 1980	\$ <u>30,000</u>
TOTAL	\$ <u>130,000</u>

All the funds provided were spent at the time of this report.

8th International Symposium on Molecular Beams
Palais des Festivals, 06400 Cannes, France
June 1-5, 1981

LBL-11373 ABS

INTERACTION POTENTIALS OF GROUND STATE XENON HALIDES

P. Casavecchia, Guozhong He, R. K. Sparks and Y. T. Lee

Materials and Molecular Research Division
Lawrence Berkeley Laboratory
and Department of Chemistry
University of California
Berkeley, California 94720 USA

For the understanding of the rare gas halide (RG-X) ground state interaction potential, and to provide useful information pertinent to RG-X lasers and halogen atom recombination studies, we undertook a few years ago several studies on the interaction potentials for the RG-F, Cl systems using the crossed molecular beams technique.¹⁻³

Recently, in order to complete the matrix of the RG-X diatomics we have investigated the interaction potentials for the systems Br + Ar, Kr, Xe and I + Kr, Xe⁴ from the measurements of total differential cross sections at several collision energies by crossing two supersonic atomic beams in a new high resolution crossed beams apparatus.⁵ Br (I) atoms were produced by thermal dissociation in a resistively heated graphite nozzle/oven⁶ using rare gas as carriers.

The scattering is described by an elastic model appropriate to Hund's case c coupling as previously done for F + Ne, Ar, Kr, Xe^{1,3} and Cl + Xe.^{2,3} The X₂¹ and I₂¹ potential energy curves arising from the ²P_{3/2} + ¹S₀ asymptote are derived by fitting calculated differential cross sections I(θ) to the data. Figure 1 shows the experimental data and the best fit calculations (solid lines) for Br-Xe. Dashed and dashed-dotted curves represent the relative contribution to I(θ) of the X₂¹ and I₂¹ potentials. The potential parameters for Br-Xe and I-Xe are listed in Table I.

The I₂¹ potential of the Br-RG and I-RG systems has a shallower well, ε, a slightly larger minimum position, r_m, and a more repulsive wall than the ¹Σ⁺ potential of the corresponding rare gas pair, while the X₂¹ potentials show a slightly stronger bonding, especially when the rare gas is xenon.

In the xenon halide series a pronounced trend on ε and r_m of the X₂¹ potential is observed going from XeI (ε 1.2 times deeper and r_m 0.1 Å shorter than for XeXe) through XeBr (ε 1.4 times deeper and r_m 0.3 Å shorter than for XeKr) and XeCl² (ε 2.1 times deeper and r_m 0.78 Å shorter than for XeAr) to XeF¹ (ε 21 times deeper and r_m 1.5 Å shorter than for XeNe). Figure 2 shows the ground state X₂¹ potential of the four xenon halides. In Fig. 3 we have reported with discontinuous lines the reduced forms of the I₂¹ potentials and with solid lines the X₂¹ potentials scaled to the corresponding I₂¹ potential. This figure illustrates the transition from van der Waals forces to almost chemical forces as we change the halogen atom. The absolute strength of the following contributions as a function of internuclear distance should contain the explanation of the V_{X₂¹}(r) bonding in RG-X: (a) less repulsion due to only a half filled p orbital along the internuclear axis, (b) a certain amount of charge transfer lessening this repulsion, (c) the effect of spin-orbit (S-O) interaction, and (d) the contribution from interatomic correlation energy (the dispersion energy in the limit of zero electron overlap). Our experimental results allow to gain some insight into their relative contribution.

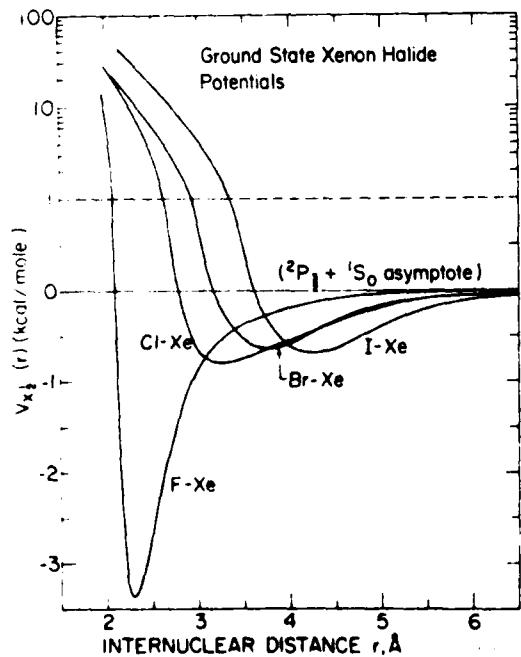
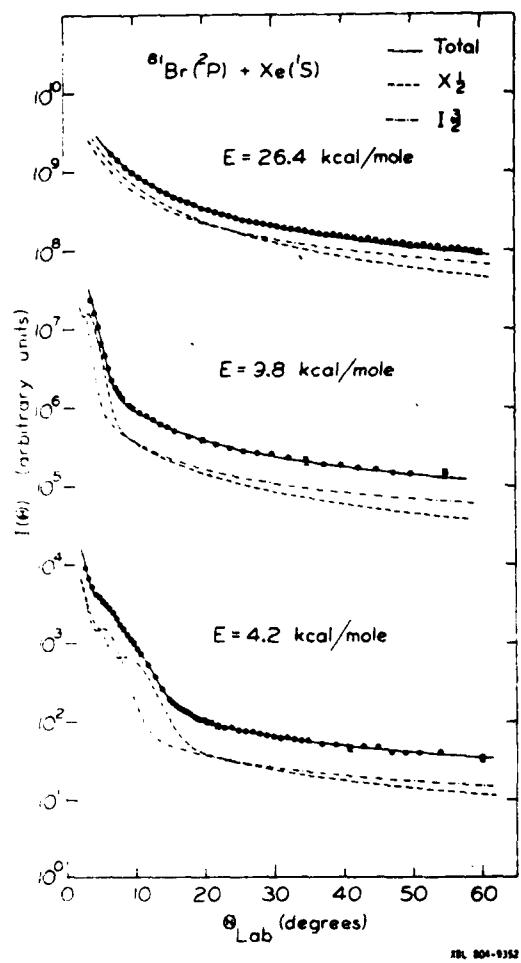


Fig. 2. Comparison of Xe-halogen interaction potentials.

TABLE I. Potential Parameters.*

System	Br-Xe		I-Xe	
	$X\frac{1}{2}$	$I\frac{3}{2}$	$X\frac{1}{2}$	$I\frac{3}{2}$
State				
ϵ (kcal/mole)	0.645	0.425	0.690	0.480
r_m (Å)	3.80	4.11	4.30	4.60
β_1	4.30	6.65	4.40	7.10
β_2	6.50	6.80	6.50	7.30
x_1	1.0405	1.0284	1.0248	1.0206
x_2	2.000	1.900	1.800	1.635
C_6 (kcal/mole·Å ⁶)	3223.	3476.	4955.	4955.
C_8 (kcal/mole·Å ⁸)	21570.	21570.	34570.	34570.

* The potential form is the MMSV function given in Ref. 1.

Fig. 1. Laboratory angular distributions of ${}^81\text{Br}$ scattered off Xe.

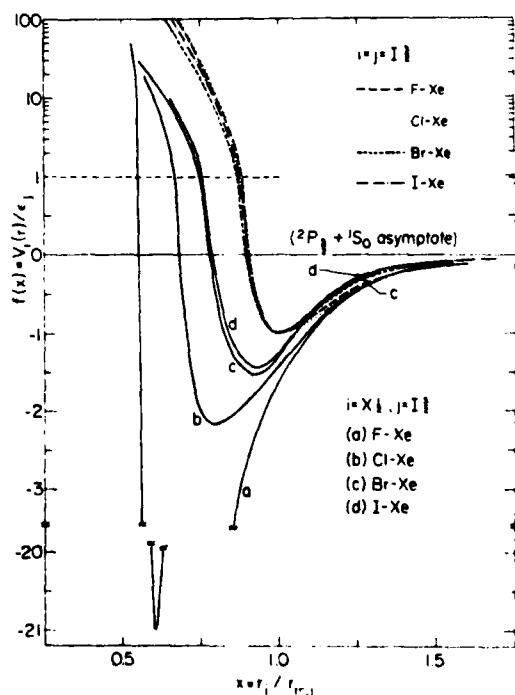


Fig. 3. Reduced potentials for Xe-halogen systems. ϵ and r_m of $I\frac{3}{2}$ potential of each of the systems was taken as unity.

The charge transfer interaction should undoubtedly play a large role and is expected to increase going from XeI to XeF because of the increase of halogen electron affinity and dissociation energy of the $\text{III}_{\frac{1}{2}}^1$ ionic state as we move towards XeF. Theory predicts a modest increase from XeI (5%) to XeF (9%).⁷ However, these predictions are somewhat doubtful in light of our results.^{1,2,4} The effect of charge transfer may account for the differences in bonding along a specific halogen-RG (RG=Ar, Kr, Xe) series (column), in line with the variation of the RG ionization potentials. But, we note that when replacing F with Cl, Br or I the relative positions of the covalent and ionic states of the separated atoms do not vary as much as when Ar is replaced by Kr or Xe, since the RG ionization potentials vary over a wider range than the electron affinity of the halogen atoms. This suggests that other factors in addition to the charge transfer play a significant role in a specific rare gas-X series (row) (i.e. Xe-X, Kr-X, where X = F, Cl, Br, I). Calculations show a dramatic increase in the admixture of π character into the ground $\text{X}_{\frac{1}{2}}^1$ state as one proceeds in the sequence XeF (0.045), XeCl (0.142), XeBr (0.379), and XeI (0.483) in connection with the increasing S-O coupling in the halogen atom.⁷ This should have a destabilizing effect on the $\text{X}_{\frac{1}{2}}^1$ state because of the repulsive character of the $\text{I}_{\frac{3}{2}}^1$ and $\text{II}_{\frac{1}{2}}^1$ (${}^2\pi_{3/2}$, ${}^2\pi_{1/2}$) states. Such an effect can be expected to be constant along a column of the RG-X matrix. The analysis of the effect of the halogen atom along a row and of the rare gas atom along a column of the RG-X matrix suggests that the interatomic correlation energy plays an important role. In particular, the explanation of the almost chemical well in XeF might also come from this important contribution. Given that the different factors contributing to the bonding are to a good degree separable, the RG-X molecules may provide insight into the nature of interatomic correlation energy in regions of significant electron overlap.

ACKNOWLEDGMENTS

This work was supported by the Office of Naval Research under Contract No. N00014-75-C-0671 and by the Division of Chemical Sciences, Office of Basic Energy Sciences, U. S. Department of Energy under Contract No. W-7405-Eng-48.

REFERENCES

1. C. H. Becker, P. Casavecchia and Y. T. Lee, J. Chem. Phys. 69, 2377 (1978); 70, 2986 (1979).
2. C. H. Becker, J. J. Valentini, P. Casavecchia, S. J. Sibener and Y. T. Lee, Chem. Phys. Lett. 61, 1 (1979).
3. C. H. Becker, P. Casavecchia, Y. T. Lee, R. E. Olson and W. A. Lester, Jr., J. Chem. Phys. 70, 5477 (1979).
4. P. Casavecchia, Guozhong He, R. K. Sparks and Y. T. Lee, to be published.
5. R. K. Sparks, Ph.D. Thesis, University of California, Berkeley, California, 1979; Y. T. Lee, J. D. McDonald, P. R. LeBreton and D. R. Herschbach, Rev. Sci. Instr. 40, 1402 (1969).
6. J. J. Valentini, M. J. Coggiola and Y. T. Lee, Rev. Sci. Instr. 48, 58 (1977).
7. P. J. Hay and T. H. Dunning, Jr., J. Chem. Phys. 69, 2209 (1978).

INTERACTION POTENTIALS FOR Br(²P) + Ar, Kr, Xe (¹S)
BY THE CROSSED MOLECULAR BEAMS METHOD

Piergiorgio Casavecchia,^a Guozhong He,^b
Randal K. Sparks,^c and Yuan T. Lee

Materials and Molecular Research Division
Lawrence Berkeley Laboratory
and
Department of Chemistry
University of California
Berkeley, California 94720

ABSTRACT

Angular distributions of Br(²P_{3/2,1/2}) scattered off Ar, Kr, and Xe (¹S₀) in the thermal energy range were measured in crossed molecular beams experiments. Interaction potentials for the X₂¹ and I₂³ states are derived by using an approximate elastic scattering analysis, which neglects interstate coupling, as previously done for F-Xe, Kr, Ar, Ne and Cl-Xe. While the Br-Xe X₂¹ potential ($\epsilon = 0.645$ kcal/mole, $r_m = 3.80 \text{ \AA}$) shows a stronger interaction than the interaction potential of Kr-Xe, the Br-Kr ($\epsilon = 0.460$ kcal/mole, $r_m = 3.90 \text{ \AA}$) and Br-Ar ($\epsilon = 0.380$ kcal/mole, $r_m = 3.73 \text{ \AA}$) X₂¹ potentials are closer to those of the corresponding rare gas pairs.

- ^a Permanent address: Dipartimento di Chimica dell'Universita, 06100 Perugia, Italy.
- ^b Permanent address: Institute of Chemical Physics, Darien, People's Republic of China.
- ^c Permanent address: Division of Chemistry and Chemical Engineering, California Institute of Technology, Pasadena, CA 91125.

The $I\frac{3}{2}$ potential for all three systems is found to have a shallower ϵ , a slightly larger r_m and a more repulsive wall than the $^1\Sigma^+$ potential of the corresponding rare gas pair. The origin of these interactions is discussed and an attempt to examine the rare gas halides ground state bonding is presented.

I. INTRODUCTION

Following the discovery of UV emission from rare gas monohalide molecules (RG-X)¹ and their subsequent use in high power UV lasers,² much work has been devoted to the spectroscopy and kinetics of these systems. The electronically excited RG-X systems have strong chemical interactions, ionic in character, quite similar to those of ground state alkali halides and have been studied extensively in both theoretical³ and experimental investigations.^{1b-d,4,5}

The picture of the interaction potentials of ground state RG-X is still not complete. In many systems, especially those involving light rare gas atoms and heavy halogen atoms, the interactions are expected to be rather weak and similar to van der Waals interactions between two rare gas atoms. On the other hand, interactions stronger than ordinary van der Waals forces recently have been shown to exist in systems such as XeF⁶⁻⁸ and XeCl.^{9,10} Since, with the exception of these two compounds, all of the observed emission spectra of the RG-X show only the diffuse structure characteristic of a bound-free transition, only very limited information on the ground state potential energy surfaces is obtained from these spectra.

The properties of the RG-X have also been the subject of intensive theoretical investigation.¹¹ Ab initio calculations have been reported on the covalent and ionic states of NeF, ArF, KrF, XeF^{3a} and XeCl, XeBr, XeI^{3b} by Dunning and Hay. An important role of these calculations has been to provide a detailed assignment of all of the features in the emission spectra of the RG-X and to predict previously

unobserved features. However, a slight rearrangement in the ordering of the excited $\text{II}_{\frac{3}{2}}^3$ and $\text{III}_{\frac{1}{2}}^1$ states was found to be necessary from the very recent detailed experimental investigations.¹² In contrast to the useful information obtained for the excited states from ab initio calculations, very little reliable information has been obtained to date on the ground state interaction. However, it is known that standard quantum chemical computational methods, even of CI type, usually fail for van der Waals type interactions.¹³

The crossed molecular beams technique has proven to be a powerful tool to give quantitative information on many adiabatic potentials not readily accessible to spectroscopic investigation.¹⁴ The extraction of interaction potentials from scattering data for systems containing non-S state open shell atoms, such as halogen (${}^2\text{P}$) or oxygen (${}^3\text{P}$) with rare gas atoms, is complicated by the fact that there is more than one potential energy surface involved. Nevertheless, it is possible to obtain meaningful results if differential cross sections are carefully measured at several collision energies covering a wide angular range, and if the nonadiabatic coupling between the relevant states is weak. Accurate potentials in the well region for F-Xe $\text{X}_{\frac{1}{2}}^1$ ⁸ and Cl-Xe $\text{X}_{\frac{1}{2}}^1$ ¹⁰ have been obtained in this laboratory from differential cross section measurements, and are in very good agreement with spectroscopically determined potentials.^{6,7,9} Also, molecular beam experiments have provided, for the first time, information on the F-Xe⁸ and Cl-Xe¹⁰ $\text{I}_{\frac{3}{2}}^3$ and $\text{II}_{\frac{1}{2}}^1$ potentials in the attractive well region. These potentials are inaccessible to spectroscopic study. The $\text{X}_{\frac{1}{2}}^1$, $\text{I}_{\frac{3}{2}}^3$ and $\text{II}_{\frac{1}{2}}^1$ potentials have also been obtained for F-Ne, F-Ar and F-Kr.¹⁵

Br-Xe was the first rare gas halide system in which laser action (at 282 nm) was achieved.^{2a} Only spontaneous emission has been reported for the Br-Kr^{4,16} and Br-Ar systems.¹⁶ Recently, from the analysis of emission spectra of Br-Ar at low pressure,¹⁷ the positions and the slopes of the repulsive part of the $X\frac{1}{2}$ and $I\frac{3}{2}$ potential curves in the Frank-Condon region have been obtained, relative to the predetermined upper state potentials. Similar information is also available for the repulsive part of the Br-Xe $X\frac{1}{2}$ potential,¹⁸ but the spectroscopic analysis has not been reported yet for Br-Kr.

In the past, the theoretical understanding of the termolecular recombination of bromine atoms in the presence of a inert gas as third body¹⁹⁻²¹ has suffered from the lack of detailed information on Br-RG interaction potentials. Knowledge of the ground state potential in RG-X systems is also important in terms of fundamental bonding theory, especially, the transition from van der Waals interaction to chemical forces. Hence, in an attempt to better understand the RG-X ground state interaction in the attractive well region, and to provide useful information pertinent to RG-Br laser and recombination studies, we have measured the differential cross sections of Br(²P) scattered off Ar, Kr and Xe(¹S) in the thermal energy range. The study reported here is a continuation of previous investigations from this laboratory on F-Ne, Ar, Kr, Xe^{8,15} and Cl-Xe.¹⁰

II. EXPERIMENTAL

The crossed molecular beams apparatus employed in this study is a newly designed higher resolution version²² of the universal crossed molecular beam machine described by Lee *et al.*²³ Supersonic beam of bromine atoms seeded in a rare gas carrier, and beams of argon, krypton or xenon are crossed at 90° in a liquid nitrogen cooled collision chamber maintained at 8×10^{-8} torr. Elastically scattered bromine atoms are detected in the plane defined by the two intersecting beams, by a triply differentially pumped rotatable ultra high vacuum quadrupole mass spectrometer detector. The laboratory scattering angle, Θ , is measured relative to the Br beam direction.

The high intensity supersonic atomic bromine beams used in this study are produced in a resistance heated graphite nozzle which has been described in detail elsewhere.²⁴ The collision energy, E, was varied by changing the carrier gas while keeping the nozzle temperature at ~1800 K. Three different seeded mixtures were generated by passing pure xenon, argon or helium through a reservoir containing liquid bromine kept at 0°C. The heated gases expand through a 0.10 mm orifice at the tip of the oven. A 0.94 mm diameter graphite skimmer was used with a nozzle-skimmer distance of 8.5 mm. The beam was collimated in a differential pumping region to an angular spread of 1.0°. The rare gas beam expanded through a 0.07 mm quartz nozzle at room temperature and was collimated to an angular divergence of 1.9°. The nozzle-skimmer distance was 5.6 mm and the skimmer diameter 0.64 mm. Under our geometrical arrangement the collision volume is always contained in the viewing angle of the detector, which has an angular resolution

of 1.0° . The velocity distributions of the beams are characterized by conventional "single shot" time-of-flight (TOF) measurements with a 30 cm flight path from the slotted disc to the ionizer. Table I gives the Br, Ar, Kr and Xe atom effective bulk flow temperature, stagnation pressure, Mach number, peak velocity and full-width-half-maximum (FWHM) relative velocity spread. The most probable collision energy for each condition is also presented. Velocity distributions of the Br seeded beams are shown in Fig. 1. The solid line is a parametric fit to the deconvoluted distribution for the Mach number and the temperature given in Table I. The performance of the graphite oven as a source of bromine atoms has been excellent, showing stability and reproducibility even after a long period (~two months) of daily operation.

Laboratory angular distributions, $I(\text{H})$, were obtained by taking from 4 to 6 scans of 20 sec. counts at each angle. Typically, signal at m/e 81 was detected. The $I(\text{H})$ are time normalized by periodically returning the detector to an arbitrary reference angle (usually 10°) in order to account for possible long term drifts in beam intensities and detector sensitivity. The rare gas target beam was modulated at 150 Hz by a tuning fork chopper. Signal plus background and background counts were obtained from a pulse counting dual scaler, synchronously gated with the tuning fork.

The graphite nozzle temperature has been estimated based on the effective temperature of the gas in the nozzle. This temperature scale is determined from analysis of the measured TOF distributions of pure helium beams produced by expansion from the graphite nozzle at several different values of the heating power (a reference thermocouple was

placed close to the tip of the nozzle). Under the normal operating temperature (~ 1800 K) a small fraction of the spin-orbit (S-O) excited $\text{Br}^{\text{2P}}_{1/2}$ atoms is produced. Because of the large S-O splitting (3685 cm^{-1}), the amount of electronic to translational relaxation during the supersonic expansion is expected to be very small. Thus, the $\text{Br}^{\text{(2P}}_{1/2})$ contribution to the beam is estimated by Boltzmann distribution and degeneracy weight to be 2.6 percent. Some small amount of undissociated bromine was also present in the beam and the small (a few percent) $^{81}\text{Br}^+$ contribution from Br_2 was taken into account from the measured angular distribution of Br_2 detected as $^{162}\text{Br}_2^+$ and the fragmentation ratio of Br_2 , $^{81}\text{Br}^+/\text{Br}_2^+$ in the ionizer. This was then subtracted from the laboratory angular distribution measured at $^{81}\text{Br}^+$ to give the final $I(\text{H})$ of $\text{Br}+\text{RG}$. In order to determine the Br_2 fragmentation ratio, the $^{81}\text{Br}^+/\text{Br}_2^+$ ratio for the ionization of Br_2 was measured as a function of oven temperature in a temperature range where negligible dissociation takes place. Almost constant values were obtained and then extrapolated to the operation temperature. A similar procedure was used satisfactorily during previous studies on F-RG^{8,15} and Cl-Xe.¹⁰

Electronic transitions are expected to occur with very small cross sections in $\text{Br}^{\text{(2P}}_J)+\text{RG}^{\text{(1S}}_0)$ collisions²⁵ in the investigated energy range, because of the large S-O splitting of the Br doublet. Thus, a negligible contribution to the measured $I(\text{H})$ is expected from inelastic events. As a consequence, no attempt was made to investigate the fine structure inelastic transitions by TOF measurements.

III. RESULTS AND ANALYSIS

Laboratory angular distributions $I(\textcircled{H})$ of Br scattered off Xe, Kr and Ar are presented on a semi-log scale in Figs. 2, 3 and 4, respectively. Exemplary error bars are shown, when visible outside the solid circle, representing ± 2 standard deviations of the mean. The nominal collision energies, E, are also indicated. While $^{81}\text{Br}^+$ was detected in the Br+Ar and Xe experiments, $^{79}\text{Br}^+$ was detected in the Br+Kr case, in order to avoid interference from the possible transmission of a small amount of $^{82}\text{Kr}^+$ at $m/e = 81$ coming from the elastically scattered krypton beam.

The $I(\textcircled{H})$ for Br-Xe (see Fig. 2) at $E=4.2$ kcal/mole clearly shows a low amplitude rainbow structure. A significant portion of rainbow on the dark side is seen at $E=9.8$ kcal/mole, while at the highest E the distribution mainly reflects the scattering from the repulsive parts of the potentials.

The $I(\textcircled{H})$ for Br-Kr (see Fig. 3) at $E=3.8$ kcal/mole shows a rainbow structure at small angles, while at $E=8.4$ kcal/mole only the dark side of the rainbow scattering is observed.

The rainbow structure is barely seen for Br-Ar at $E=2.9$ kcal/mole shown in Fig. 4. A slower Br beam would have been desirable in order to obtain a significantly lower collision energy and consequently shift the rainbow oscillation to larger angles. Unfortunately, the nozzle temperature cannot be decreased without a drastic reduction of Br_2 dissociation. Also, Br_2 was already seeded in the heaviest rare gas, Xe. The unusual shape of the $I(\textcircled{H})$ at large angles for Br-Ar, particularly pronounced at high E, is due to the detection of heavier

atoms in this system. When one detects the heavy particle, particles scattered at two different center-of-mass (CM) angles are observed at the same laboratory angle. The heavier bromine atom is kinematically constrained to scatter within a lab angle of approximately 38° at $E=14.8$ kcal/mole and of approximately 55° at $E=2.9$ kcal/mole, as predicted from the most probable Newton diagrams. As we approach the edge of the elastic Newton circle, the differential cross section shows a rather broad peak in the vicinity of these cutoff angles due to the nature of the transformation Jacobian which relates the lab and CM reference frames.¹⁴ The angular distributions shown in Fig. 4 have been truncated at large angles near the onset of this peak. This effect has been properly accounted for in the data analysis.

The procedure of analysis to obtain interaction potentials for the states involved in the scattering by fitting the $I(H)$ is the same as that of Refs. 8, 10 and 15 and will only be briefly discussed here. The molecular electronic states arising from the four-fold degenerate ground state $Br(^2P_{3/2}) + RG(^1S_0)$ asymptote are the doubly degenerate X_Σ^1 (or I_Σ^1) and I_Σ^3 states, in Hund's case c notation. (In Hund's case b and a notation these are designated $^2\Sigma_{1/2}^+$ and $^2\Pi_{3/2}$, respectively). From the doubly degenerate spin-orbit excited manifold, $^2P_{1/2} + ^1S_0$, one obtains the II_Σ^1 state ($^2\Pi_{1/2}$ in Hund's case a). In this notation the $1/2$ or $3/2$ represents the Ω quantum number (the projection of the total electronic angular momentum along the internuclear axis); the X, I, II represents the ordering of the states.

It has been shown that inelastic cross sections are much smaller than the elastic cross sections for F-Ar, Xe and Cl-Xe in the thermal

energy range.²⁵ Since the inelastic cross sections should be even smaller for the Br-RG systems, the total differential cross section can be written as a sum of elastic differential cross sections $\sigma_{X_{1/2}}(\theta)$, $\sigma_{I_{3/2}}(\theta)$, and $\sigma_{II_{1/2}}(\theta)$ for the three states X 1/2, I 3/2 and II 1/2. Each state's contribution is given its appropriate statistical weight and added to give a total $\sigma_T(\theta)$:

$$\begin{aligned}\sigma_T(\theta) &= \frac{1}{2} (0.974) [\sigma_{X_{1/2}}(\theta) + \sigma_{I_{3/2}}(\theta)] + (0.026)\sigma_{II_{1/2}}(\theta) \\ &= 0.487 \sigma_{X_{1/2}}(\theta) + 0.513 \sigma_{I_{3/2}}(\theta)\end{aligned}\quad (1)$$

The approximation $\sigma_{II_{1/2}} = \sigma_{I_{3/2}}$ has been used. Because of the small contribution of $\sigma_{II_{1/2}}$ to σ_T , this approximation should have a negligible effect in the evaluation of the $V_{X_{1/2}}(r)$ and $V_{I_{3/2}}(r)$. Each elastic differential cross section in the CM system is calculated independently by partial wave analysis with JWKB phase shifts²⁶ to give a calculated $I(\textcircled{H})$ with appropriate averaging over the velocity and angular distributions of the two beams and the detector acceptance angle. Comparison of calculated $I(\textcircled{H})$ with experimental $I(\textcircled{H})$ provides the basis for evaluation of the interaction potentials.

The accuracy of this case c elastic approximation has been supported in two ways as shown in previous studies.^{8,10,15} First, in the F-Xe case, where an accurate spectroscopically (RKR) determined potential was available,^{6,7} for the $V_{X_{1/2}}(r)$ this scattering method was found to give results in good agreement.⁸ Scattering studies¹⁰ were also fruitful in corroborating a spectroscopically obtained

$v_{X\ 1/2}(r)$ for Cl-Xe.⁹ Second, rigorous coupled-channel scattering computations²⁵ show that the elastic approximation reproduces all of the gross features of the experimental $I(H)$ very well for F-Xe, F-Ar and Cl-Xe at thermal energies. The approximation is expected to hold even better for systems where the open shell atom has larger S-0 splitting, such as bromine and iodine.

The physical meaning of the elastic analysis used here is that the electron orbital and spin angular momentum remain coupled throughout the collision and coupling is weak between the adiabatic states of the total electronic Hamiltonian including S-0 interaction. Hund's case c affords a good description of this situation, with Ω being the good quantum number.^{27,28} This model corresponds to describing the collisions by elastic scattering separately on each of the adiabatic potential energy curves. The same model has been followed by Aquilanti *et al.*²⁹ in the analysis of their results on the scattering of magnetically selected $O(^3P)$ atoms with rare gases at thermal energies; absolute integral cross sections were measured. Our model is physically very different from the elastic models (see Ref. 30 and references cited therein) previously used to describe thermal collision of systems with small S-0 coupling, such as $Li(^2P)-RG$. Those models are based on the assumption that, during the collision, the electron orbital angular momentum and spin remain uncoupled, and that coupling is weak between the adiabatic states of the electrostatic electronic Hamiltonian.

In the analysis of the experimental results of Figs. 2, 3 and 4, the interaction potentials $V(r)$ are chosen to be the flexible, piece-wise analytic, Morse-Morse-switching function-van der Waals (MMSV) form. The MMSV reduced form is given by:

$$f(x) = V(r)/\epsilon$$

$$x = r/r_m$$

$$f(x) = \exp(2\beta_1(1-x)) - 2\exp(\beta_1(1-x)), \quad 0 < x \leq 1$$

$$= \exp(2\beta_2(1-x)) - 2\exp(\beta_2(1-x)) = M_2(x) \quad 1 < x \leq x_1$$

$$= SW(x) M_2(x) + (1 - SW(x)) W(x) \quad x_1 < x < x_2$$

$$= -C_{6r}x^{-6} - C_{8r}x^{-8} = W(x) \quad x_2 \leq x < \infty$$

and

$$SW(x) = \frac{1}{2} \left[\cos \frac{\pi(x - x_1)}{(x_2 - x_1)} + 1 \right],$$

where

$$C_{6r} = \frac{C_6}{(\epsilon r_m^6)}, \text{ and } C_{8r} = \frac{C_8}{(\epsilon r_m^8)};$$

ϵ and r_m are the depth and position of potential minimum.

The van der Waals C_6 constant is estimated by the Slater-Kirkwood formula³¹ for effective number of electrons:

$$C_6 = \frac{3}{2} \frac{\alpha(Br) \alpha(RG)}{[\alpha(Br)/N(Br)]^{1/2} + [\alpha(RG)/N(RG)]^{1/2}}$$

where $\alpha(Br)$ and $\alpha(RG)$ are the polarizabilities of bromine atoms³² and rare gas atoms (RG=Ar, Kr, Xe),³³ respectively, and N is the effective number of electrons [N(Br)=5.7, N(Ar)=6.0, N(Kr)=6.7 and N(Xe)=7.93]. The halogen atoms are known to have an anisotropy of polarizability, being $\alpha(m_j=\pm 1) > \alpha(m_j=0)$.^{34,35} The anisotropy of Br, for which no experimental or theoretical value was found, has been

taken to be equal to that of Cl (~15%),³⁵ which is slightly larger than that of F (~12%).³⁴ The C_6 constant of I 3/2 is calculated to be larger than X 1/2 state, reflecting the larger polarizability of the II symmetry. This implies a slow curve crossing for the X 1/2 and I 3/2 states at fairly large r , as obtained for the F-RG and Cl-Xe. The C_6 for the X 1/2 state has the long range mixture of one-third $^2\Pi$ and two-thirds $^2\Sigma^+$ character,^{3a-b,36} while I 3/2 is purely $^2\Pi$. The II 1/2 state asymptotically approaches two-thirds $^2\Pi$ and one-third $^2\Sigma^+$, but we have used the C_6 of the I 3/2 state, since we have assumed $V_{II\ 1/2} = V_{I\ 3/2}$. This should have a negligible effect. The C_8 constants are estimated from the Kr-Rg (Rg=Ar,Kr,Xe) C_8 constants.³³ The permanent quadrupole induced dipole induction constant, varying as R^{-8} at long-range, contributes only a small fraction of the C_8 dispersion term, as estimated from the permanent quadrupole moment of Br³⁷ and the polarizability of the rare gases.³³ It has thus been neglected, as have higher order constants, because of their uncertainty and very small contribution.

Initially, the I 3/2 potentials for all three systems were assumed to be very near the corresponding one electron richer rare gas pair Kr-RG. Justification of this choice is based on the closed shell-closed shell electronic configuration for the II symmetry, having the fully occupied bromine p orbital along the internuclear axis.^{3a} With this assumption the parameters ϵ , r_m , β_1 , and β_2 for the X 1/2 potential were varied in an attempt to match the calculated and experimental $I(\text{H})$ at three collision energies for Br-Xe and two collision energies for Br-Kr and Br-Ar. In order to reach a good simultaneous

fit in all the angular ranges, the original estimate of the $I_{3/2}$ potential needed to be slightly modified in the well depth for all three systems. In the Br-Ar case, r_m also needed to be modified. A little variation was also made in the x_1 and x_2 values for both potentials. The best fit $I(\textcircled{H})$ are reported as a solid line in Figs. 2, 3 and 4 for Br-Xe, Br-Kr and Br-Ar, respectively. The calculated $I(\textcircled{H})$ are scaled to the data by a constant scaling factor which is determined by the minimization of a chi(χ)-square goodness of fit measure. The derived $V_{X\ 1/2}(r)$ and $V_{I\ 3/2}(r)$ potentials are depicted in Figs. 5, 6 and 7. The resulting potential parameters are listed in Table II, where those of the Kr-Ar,³⁸ Kr-Kr³⁹ and Kr-Xe⁴⁰ systems are also reported for comparison. The contribution to $I(\textcircled{H})$ by each of the potentials reported in Table II is shown with discontinuous lines in Fig. 2 for Br-Xe at three energies and in Figs. 3 and 4 for Br-Kr and Br-Ar only at the lowest E. The relative weights from Eq. (1) are used in the $I(\textcircled{H})$ plot for all energies.

Measurements of $I(\textcircled{H})$ at more than one collision energy and covering a wide angular range allow an accurate estimate of the two interaction potentials which contribute most to the scattering. In fact, rainbow and supernumerary rainbow positions and relative intensities are very sensitive to potential well depths and curvature.^{14,41} Also, the range parameters of the potentials are sensitive to the ratio of rainbow to wide angle scattering intensities. In particular, the relative magnitude of the cross section at large angles to that at the rainbows gives a good indication of the "size" of the atoms, that is, the position of the repulsive wall of the

potential. Quite strong sensitivity to the $V_X 1/2$ well of Br-Xe was found in the lowest energy angular distributions. As shown in Fig. 2 at $E=4.2$ kcal/mole, the main rainbow and the first supernumerary from the $X 1/2$ potential are observed (dashed line). The main rainbow produced by the $I 3/2$ potential is also clearly observed (dot-dashed line) providing quite good sensitivity of $I(H)$ also to the $V_I 3/2$. A pronounced quenching can be noted when the two separate contributions are added according to Eq. (1) to give the best fit $I(H)$; but, two major low amplitude oscillations can still be seen. At $E=9.8$ kcal/mole significant contribution from the $V_X 1/2$ rainbow is still observed. While at the highest E , the separated contributions clearly show that scattering is mainly taking place from the repulsive walls of the two potentials. The data presented in Fig. 2 allows an estimate of ϵ and r_m parameters for the $V_X 1/2(r)$ to within ± 5 percent and for the $V_I 3/2(r)$ to within ± 7 percent for Br-Xe. Estimates of the uncertainties were obtained by making a matrix of plots of computed $I(H)$, with the different matrix elements corresponding to potentials with differing (ϵ , r_m) values (always keeping the overall potential's to a reasonable shape). The ϵ and r_m values were varied for the two different states until the fits to the experimental $I(H)$ became poor. These computations were made for every system at all the energies investigated. Possible error in the Morse β parameters are likely to be of similar magnitude, based on their observed influence on the $I(H)$ during the trial and error fitting procedure. Because the quality of the data at large angles are also quite good, fairly high sensitivity to the repulsive walls is obtained. Of course, these

potentials for Br-Xe can be considered valid on the repulsive wall only to about 26 kcal/mole, which is the highest collision energy studied.

Figure 3 shows that, for the Br-Kr system at E=3.8 kcal/mole, the main rainbow oscillation from both potentials is observed. Again, the total $I(\text{H})$ is strongly quenched, particularly because the oscillatory behaviors of the two different rainbow patterns are out of phase.

Figure 4 shows that the rainbow structure from the $V_X \frac{1}{2}$ is almost fully resolved at E=2.9 kcal/mole for Br-Ar, while only a portion of the dark side of the rainbow from the $V_I \frac{3}{2}$ is seen. For Br-Kr and Br-Ar, a slightly lower sensitivity of $I(\text{H})$ to the ϵ and r_m parameters of the $V_X \frac{1}{2}$ and $V_I \frac{3}{2}$ than for Br-Xe is observed. The estimated maximum uncertainties in ϵ and r_m , obtained with the procedure described above, for both systems are within ± 7 percent and ± 10 percent for $V_X \frac{1}{2}(r)$ and $V_I \frac{3}{2}(r)$, respectively. Uncertainties for the Morse parameters, which govern the shape of the potentials near the minimum and the repulsive wall are similar to those obtained for ϵ and r_m .

We note that the final x_2 joining points obtained are slightly larger than the cut-off points $2 [\langle r_{Br}^2 \rangle^{1/2} + \langle r_{RG}^2 \rangle^{1/2}]$ suggested⁴² for the validity of the power series expansion, which were taken as initial guesses. The C_6 and C_8 van der Waals constants were kept fixed during the fitting, because the differential cross section is not very sensitive to the long range part of the potential. Consequently, no accurate information is gained on the long range van der Waals attraction.

IV. DISCUSSION

A detailed characterization of the bound-free emission spectra of the RG-X would require a knowledge of the potential energy curves of both the upper and lower states over the relevant range of energy and internuclear distance, in addition to the variation of the electronic transition moment with internuclear distance and the vibrational distribution of the emitting states. The close analogy of the excited RG-X with ground state alkali halides has provided the basis for the interpretation and the prediction of the emission spectra.⁴ From trial and error semiclassical simulation of the diffuse spectra, information on the upper states and approximate shapes for the lower state potential curves in the Frank-Condon region have been obtained.^{17,18} But, no unique potential curves can be derived from diffuse emission spectra alone. In fact, the form of a spectrum is determined principally by the relative shapes of the upper and lower potentials.

A. Br-Xe.

The Br-Xe emission spectrum has been studied by several groups.^{4,5,1c} Crude estimates of the upper III 1/2 (B) state vibrational frequency, ω_e' , were given. The most recent and detailed investigation was performed by Tellinghuisen et al.,¹⁸ who attempted a quantitative analysis of the high pressure III 1/2 \rightarrow X 1/2 (B \rightarrow X) bands. The ratios of ω_e' 's and the dissociation energies, D_e' 's, between XeBr and CsBr are found to be 0.80 and 0.91, respectively. In general ω_e' and D_e' values are always smaller for the RG-X, than for alkali halides, in qualitative accord with the replacement of

the alkali ion by the larger and softer rare gas ion. But, no simple quantitative explanation for the trend observed along the series has been found. The R_e' value of 2.96 Å used for the III 1/2 state in the spectrum simulation¹⁸ is 0.11 Å shorter than R_e' for CsBr⁴³ and is significantly shorter than the theoretically calculated value of 3.38 Å.^{3b} The ground state X 1/2 potential curve was found to have an appreciable slope in the Frank-Condon region, $-(dV_{X\ 1/2}/dR)_{R_e'} = 5.8 \pm 0.6 \text{ kcal/mole/Å}$.¹⁸ The absolute $V_{X\ 1/2}(R_e')$ was estimated to be $1.4 \pm 0.6 \text{ kcal/mole}$. The repulsive wall of Br-Xe $V_{X\ 1/2}(r)$ obtained from the present study gives $-(dV_{X\ 1/2}/dR)_{R_e'} = 2.96 \text{ Å}^{5.6-0.8} \text{ kcal/moleÅ}$ and $V_{X\ 1/2}(R_e') = 1.0 \pm 0.3 \text{ kcal/mole}$ (see Fig. 5 and Table II). These appear to be in good accord with the spectroscopic estimates.¹⁸ Thus, our results also represent indirectly a corroboration of the estimate of Tellinghuisen *et al.*¹⁸ of the shape and position of the upper state potential, in addition to providing a more complete picture of $V_{X\ 1/2}(r)$ and $V_{I\ 3/2}(r)$. Because of the large S-O interaction in Br(²P) the non-ionic manifold should be clearly split into three distinct states.^{3b} This produces important effects as far as the emission properties are concerned.¹¹ But, there appears to be no experimental observation as yet of the theoretically predicted band ending on the I 3/2 state. The III 1/2 \rightarrow II 1/2 broad band has been investigated very recently.⁴⁴ The theoretical curves of Ref. 3b were used in an attempt to simulate the temperature dependence of the unstructured experimental spectra. Consequently, very little information exists on the $V_{I\ 3/2}(r)$ and $V_{II\ 1/2}(r)$. No direct information

is obtained about the latter from our study. However, in general, $V_{II\ 1/2}$ can be derived, as shown in Ref. 25, from the $V_{X\ 1/2}$ and $V_{I\ 3/2}$.

Our results for the Br-Xe system show that the $V_{X\ 1/2}(r)$ has a stronger interaction than the $V_X' \Sigma^+$ of Kr-Xe (see Table II).⁴⁰ The well depth is about 0.18 kcal/mole (~40 percent) deeper and r_m 0.32 Å shorter. Also, the Br-Xe inner wall appears to be much less repulsive. The higher polarizability of Br with respect to Kr cannot alone explain these effects. The binding energy of the $X\ 1/2$ state decreases in the sequence F-Xe, Cl-Xe and Br-Xe, probably in relation to the smaller electronegativity of the heavier halogens. In the F-Xe bonding^{6-8,38} chemical forces seem to be operative; on the other hand, for Br-Xe, although the binding energy is closer to that of Kr-Xe, it does not seem to be entirely due to van der Waals interaction. The following contributions as a function of internuclear distance might provide the explanation of the $V_{X\ 1/2}(r)$ bonding in RG-X: (a) less repulsion due to an only half filled p orbital along the internuclear axis, (b) a certain amount of charge transfer further lessening this repulsion, and (c) the contribution from interatomic correlation energy (the dispersion energy in the limit of zero electron overlap). These arguments are used also by Krauss and Liu⁴⁵ to give a plausible explanation of the F-Xe $V_{X\ 1/2}(r)$. The S-O interaction was not included in their calculation, nor the variation of the intra-atomic correlation and the coupling between the inter- and intra-atomic correlation energy. The latter are difficult problems when one wishes to perform accurate computations.^{46,13}

Hay and Dunning^{3b} estimated the ionic contribution in configuration interaction (CI) to the X 1/2 state, at an interatomic distance which corresponds to the calculated R_e' of the ionic state. Ionic contribution was found to decrease modestly from F-Xe (9.4 percent) to Br-Xe (5.2 percent). The interaction between the covalent and ionic states provides a stabilizing ion-pair contribution to the X 1/2 state well depth, and a repulsive contribution to the III 1/2 curve. The results of our studies^{8,10,15,47} appear to support a substantial charge-transfer in the X-Xe X 1/2 states, which does not seem to vary smoothly going from F-Xe to Br-Xe. As a consequence, the question of the accuracy of the CI calculations is raised. These calculations show the III 1/2 curve to lie above the II 3/2 curve only at internuclear distances shorter than R_e' of the III 1/2 excited state.^{3a-b}

However, recent experiments¹² have confirmed earlier indications⁴⁸ of a reversed ordering of the III 1/2 and II 3/2 states in xenon halides. The II 3/2 state is found to lie below the III 1/2 state by about 600 cm^{-1} , 128 cm^{-1} and 80 cm^{-1} in F-Xe, Cl-Xe and Br-Xe, respectively.¹² Krauss^{3c} points out that the energy splitting between the ionic III 1/2 and II 3/2 curves is very dependent on the configuration mixing term between the ground X 1/2 state and the excited III 1/2 state. He notes also that the degree of mixing has little effect on the lifetime of the emitting states, which are indeed rather well predicted by the theory.¹¹ However, the possibility that CI calculations do not properly account for configuration interaction with the charge transfer state is also recognized by Hay and Dunning.^{3b} From a comparison of calculated w_e' , D_e' and R_e'

with experimentally derived values, a fairly good agreement is observed between values of ω_e' , while the calculated values of R_e' are systematically larger by 0.2-0.4 Å and the values of D_e' systematically lower.¹¹ The theoretical values were reasoned to be more reliable than the experimental R_e' values.^{3b} However, the R_e' derived from experiments for the III 1/2 state of F-Xe^{6,7} and Cl-Xe,⁹ where rovibrational and vibrational structure, respectively, is resolved, are expected to be reliable. The results obtained for Br-Xe from our experiments also seem to confirm the "spectroscopic" value rather than theoretical calculations. The longer internuclear distance for ground state (X 1/2) relative to excited state (III 1/2) XeBr with respect to XeF and XeCl explains the lower absorption intensity for XeBr observed in spectroscopic studies of matrix-isolated XeF, XeCl and XeBr.⁴⁹

While the Br-Xe $V_X 1/2$ clearly deviates from the corresponding Kr-Xe potential, as discussed above, the $V_I 3/2$ shows a close resemblance to it, being just 0.04 kcal/mole shallower and slightly more repulsive (see Table II). Since the polarizability of Br in the I 3/2 state of XeBr is higher than that of Kr in the Kr-Xe $X^1\Sigma^+$ state, the result has to be interpreted as a manifestation of a more repulsive character of the II type interaction in the open shell-closed shell system, with respect to the Σ type closed shell-closed shell pure van der Waals interaction in the rare gas pair. This is due to the fully occupied p orbital along the molecular axis and the larger $\langle \bar{r}^2 \rangle$ of the outermost electron cloud of Br. The ionic-covalent configuration mixing is calculated to be much smaller for the I 3/2 state than

for the $X\ 1/2$ state.^{3c} This is reflected in its significantly less bound character and in the smaller $II\ 3/2 \rightarrow I\ 3/2$ transition moment.

The calculated $X\ 1/2$ and $I\ 3/2$ curves of Br-Xe are found to be essentially repulsive^{3b} and consequently do not compare satisfactorily with the curves derived here. Nevertheless, ab initio calculations undoubtedly are of more value in describing the repulsive walls fairly far away from the region of the wells than are these scattering results, provided that the spin-orbit energy remains approximately constant with the internuclear distance.

B. Br-Kr,Ar

As can be seen from Table II, the results obtained for the Br-Kr and Br-Ar systems show that the well depth of the $V_{X\ 1/2}$ is only slightly deeper than those in the corresponding rare gas pairs, namely Kr-Kr³⁹ and Kr-Ar.³⁸ It seems that they begin to manifest a clear analogy to the rare gas pair. In fact, the slightly deeper ϵ (about 0.06 and 0.04 kcal/mole for Br-Kr and Br-Ar, respectively), and the modestly shorter r_m (about 0.1 Å for both systems) can be reasonably associated with the slight differences in the size of the outermost electron cloud between the bromine and krypton atoms and their polarizabilities anisotropic in the case of Br. Kr has a smaller radius for its outermost electrons since the Kr nucleus has one more nuclear charge than Br.³⁷ But Br has a larger polarizability than Kr, which will tend to offset the effects of the difference on atomic sizes. One should expect that in the $X\ 1/2$ state of Br-Kr and Br-Ar the bromine atom, having only one 4p electron on axis with the rare gas atom, could approach closer and the binding energy should be higher and the minimum position shorter than in Kr-Kr and Kr-Ar.

The experimentally derived I 3/2 potentials for Br-Kr and Br-Ar (see Table II and Figs. 6 and 7) have a larger r_m (by ~0.1 Å), a shallower ϵ (by about 0.05 and 0.08 kcal/mole for Br-Kr and Br-Ar, respectively), and are slightly more repulsive than the $X^1\Sigma^+$ potentials of Kr-Kr and Kr-Ar. A similar situation was observed also for the V_I 3/2 of Br-Xe, and as discussed in Sect. IV-A. The binding energy of the I 3/2 state of the rare gas bromides, being less than the rare gas-rare gas systems, indicate the larger Br outermost electron radius and the fully occupied p orbitals contribute to repulsion at larger values of internuclear separation.

Since Br has a larger polarizability than Kr, the short range repulsion seems to dominate the difference between V_X 1/2 and V_I 3/2 in the Br-RG (RG=Ar,Kr and Xe) diatomics. In an attempt to better understand the bonding, on the basis of our results we can try to estimate the relative contribution of the effect (i) of the outermost electron radius $\langle \bar{r}^2 \rangle$, (ii) of the angular distribution of the outermost orbitals (i.e., half or fully occupied p orbital), and (iii) of the charge transfer. For the I 3/2 state we can neglect the effect of a poorly known contribution of charge transfer and the effect of the half-filled orbital. We can explain, then, the bonding of the I 3/2 state of the RG-Br to be mainly determined by the size of the outermost electron cloud of the halogen atom, when compared with the $X^1\Sigma^+$ state of the corresponding rare gas pair. This should also be valid for the I 3/2 state of the other RG-X. While for the X 1/2 state all the factors mentioned above give some contribution to the bonding. In this case, there is not sufficient information to allow us to distinguish

their relative contribution. However, we can gain some insight into this question on the basis of the following observations.

Going from Br-Xe to Br-Kr and Br-Ar, only changing the nature of the rare gas, two factors can explain the trend of the X 1/2 potentials: the more compact size and consequently lower polarizability of Kr and Ar with respect to Xe; and their higher ionization potential which pushes the energy of the ionic states higher. This last factor diminishes the degree of ionic-covalent mixing,^{3c} which is one of the important contributing factors to the ground state X 1/2 binding energy. As can be seen from Figs. 5-7 and Table II, the X 1/2 state of Br-Xe has a deeper well and a significantly shorter r_m than Br-Kr and Br-Ar with respect to the $X^+ \Sigma^+$ state of the corresponding rare gas pair; it is also less repulsive. Hence, the repulsive walls of the X 1/2 and I 3/2 potentials are more clearly distinct in the Br-Xe case. The above observation suggests that the effect of charge-transfer may account for the differences in bonding along a specific halogen-RG series (i.e., Br-RG, Cl-RG, and F-RG, where RG = Xe, Kr, and Ar). The replacement of Xe with Kr or Ar was shown to have a dramatic effect on the rare gas monofluorides.¹⁵ In this case the nature of the halogen, the strongly electronegative fluorine atom, also plays an important role in the anomalous trend within the series. In particular, the explanation of the almost chemical well in the X 1/2 state of F-Xe^{6-8,15,45} should also depend critically on the contribution of the interatomic correlation energy. In fact, we note that when replacing F with Br, the relative positions of the covalent and ionic states of the separated atoms do not vary as much as when Ar is

replaced with Xe, since the rare gas ionization potentials vary over a much more extended range than the electron affinities of the halogen atoms. Since the $V_X^{1/2}$ well depth and equilibrium distance vary quite dramatically as we move from XeF ($\epsilon = 3.359$ kcal/mole, $r_m = 2.293 \text{ \AA}$) through XeCl ($\epsilon = 0.80$ kcal/mole, $r_m = 3.23 \text{ \AA}$) to XeBr ($\epsilon = 0.645$ kcal/mole, $r_m = 3.80 \text{ \AA}$), this seems to suggest that other factors in addition to the charge-transfer mixing play a significant role in a specific rare gas-X series (i.e., Xe-X, Kr-X, where X = F, Cl and Br). One of these factors may be the effect of the half occupied p orbital, "hard" in the small fluorine atom, "soft" in the large bromine atom. Another, very important factor,⁴⁵ is again the effect of the interatomic correlation energy. Now, from a comparison of F-Kr¹⁵ with Br-Kr (or F-Ar¹⁵ with Br-Ar) and F-Xe⁸ with Br-Xe, we are lead to conclude that F-Xe is quite a "peculiar" system. It would seem that Xe is chemically more different from Kr or Ar, than F is from Cl or Br. It should be noted that in the F-Xe system we have a limiting situation: the smallest and most electronegative halogen atom with the most polarizable and consequently most easily ionizable rare gas atom. This situation may produce, as the result of several additive effects, a significant negative energy contribution. What we want to emphasize is that the presence of Xe is more critical than that of F, since, in the Kr-X or Ar-X (X=F, Cl, Br) series we don't observe such a "peculiar" system as F-Xe.

In the RG-X series the S-O interaction may also play an important role in the trends of the X^{1/2} state bonding. Calculations show a dramatic increase in the admixture of π character into the ground

X 1/2 state as one proceeds in the sequence XeF (0.045) XeCl (0.142), XeBr (0.379), and XeI (0.483).^{3b} This should have a destabilizing effect on the X 1/2 state, because of the repulsive character of the I 3/2 and II 1/2 states. Such an effect can be expected to have the same importance in a specific halogen-RG series (RG=Ar,Kr and Xe). However, in their model Hay and Dunning neglect any interaction between covalent and ionic states in determining the S-O coupling elements. Little is known about the S-O interaction for open shell molecules. Other factors, such as the inclusion of a relativistic core and the influence of relaxing the restriction of the frozen Hartree-Fock core in Xe, have been shown to have negligible effect on calculated curves.^{3b} Given that the different factors contributing to the bonding are to a good degree separable, the RG-X molecules may provide insight into the nature of interatomic correlation in regions of significant electron overlap.

It will be interesting to complete the series of the RG-X molecules, by studying also the RG-I systems.⁵⁰ It will then be possible to examine in a more complete fashion the effect of the halogen atom and/or of the rare gas atom on the potential energy curves of the RG-X.

No theoretical studies have been performed on the Br-Kr and Br-Ar systems. Recently, a systematic study of the Br-Ar emission spectra has been published by Golde and Kvaran.¹⁷ By using spectroscopic and kinetic data, valuable information was obtained on the potential curves and vibrational distributions of the emitting states. Three partially overlapping transitions, ending on the three states of the lower Br-Ar

manifold, contribute to the continuum spectra, with obvious complication for the analysis. The upper III 1/2 and II 3/2 emitting states were predetermined by analogy with KBr and were taken to be essentially identical. The slopes and the absolute energy of the $V_X 1/2$ and $V_I 3/2$ in the Frank-Condon region were estimated. The most precise results were obtained for the $V_X 1/2$ from the analysis of the highly structured III 1/2 \rightarrow X 1/2 continuum. The following estimate was determined:

$$-(dV_{X 1/2}/dR)_{R_e = 3.09 \text{ \AA}} = 13.1 \text{ kcal/(mole \AA)} \text{ and}$$

$V_{X 1/2}(R_e = 3.09 \text{ \AA}) = 4.6 \text{ kcal/mole}$, with an uncertainty of about 10 percent. Our best fit results give:

$$-(dV_{X 1/2}/dR)_{R_e = 3.09 \text{ \AA}} \approx 12.6 \text{ kcal/(mole \AA)} \text{ and}$$

$$V_{X 1/2}(R_e = 3.09 \text{ \AA}) \approx 1.7 \text{ kcal/mole.}$$

Less precise information was obtained by Golde and Kvaran¹⁷ on the I 3/2 curve, because of unseparable emission overlaps in the II 3/2 \rightarrow I 3/2 system. The I 3/2 curve was found to be slightly more repulsive than the X 1/2, the best fit values are:

$$-(dV_{I 3/2}/dR)_{R_e = 3.09 \text{ \AA}} \approx 29 \text{ kcal/mole/\AA}$$

and

$$V_{I\ 3/2}(R_e' = 3.09 \text{ \AA}) \cong 7.4 \text{ kcal/mole}.$$

Our best fit results give:

$$-(dV_{I\ 3/2}/dR)_{R_e'} = 28 \text{ kcal/mole/\AA}$$

and

$$V_{I\ 3/2}(R_e' = 3.09) \cong 4.8 \text{ kcal/mole} .$$

Thus, a comparison of the repulsive walls for the $V_{X\ 1/2}$ and $V_{I\ 3/2}$ reported in Table II and depicted in Fig. 7 with those determined from the spectroscopic study¹⁷ shows good accord for the slopes of the curves and a satisfactory accord for the absolute positions. Although the spectroscopic estimate of $V_{X\ 1/2, I\ 3/2}(R_e')$ falls within the limits of our uncertainties (see Sect. III) for the Br-Ar potentials, we note that a perfect accord would be obtained if R_e' was taken to be 2.93 Å, which is 0.11 Å larger than R_e of KBr (2.82 Å).

An interesting result of Ref. 17 was the need to invoke van der Waals minima in the X 1/2 and I 3/2 states of Br-Ar in order to satisfy the upper limit for the electronic energy of the III 1/2 (B) state, as derived from the temperature dependence of the spectra. Our work provides also the necessary information on the relative energies of the X 1/2 and I 3/2 states in the well region. As discussed by

Golde and Kvaran, another way to satisfy the upper limit of the III 1/2 state is to lower the energy of it by changing R_e' towards smaller values, as suggested by our results. The emission spectra of Br-Kr has also been recorded, but no analysis has been yet reported.^{4,16} Similar features as for Br-Ar were observed.

C. Recombination Studies.

Among the termolecular recombination processes of halogen atoms in the presence of a third body, those of bromine atoms in a rare gas environment have been studied most extensively, both experimentally and theoretically.⁵¹ In general, the lack of detailed information on the halogen-rare gas interaction potentials made it difficult to evaluate the theoretical understanding of the termolecular recombination processes.¹⁹⁻²¹

Clarke and Burns^{19b} computed rate constants for the recombination of bromine atoms in the presence of He, Ar and Xe by 3-D classical trajectory methods. They found that the radical molecules complex (RMC) mechanism was dominant for Ar and Xe at low temperatures; fairly good agreement between experiment and theory was obtained. A critical parameter in their model was the well depth of the Br-RG potential. Since the equilibrium constant for the process $\text{Br} + 2 \text{RG} \rightleftharpoons \text{Br-RG} + \text{RG}$ increases approximately exponentially with $\epsilon_{\text{Br-RG}}$, the rate constant for the RMC recombination was significantly affected by $\epsilon_{\text{Br-RG}}$. In their study, the interaction potential between Br and RG was assumed to be of the Lennard-Jones (L-J) form with the following parameters: RG=Ar, $\epsilon=1.0$ kcal/mole, $\sigma=3.0$ Å; RG=Xe, $\epsilon=1.0$ kcal/mole, $\sigma=3.5$ Å. A slightly shallower Br-Ar potential represented by a Morse function

($D_e=0.865$ kcal/mole, $r_e=3.51$ Å) was used in another calculation.^{19c} When the parameters of the Br-Ar potential were treated as adjustable constants in the phase space theory,^{19c} better agreement with experiment was obtained. The optimum parameters were $D_e=1.26$ kcal/mole, $r_e=2.4$ Å. No reliable information was available ther. to indicate which set of parameters would better represent the Br-Ar potential.

The ϵ_{Br-Ar} was lowered to 0.5 kcal/mole in a study^{19e} where the recombination path via unstable Br-Ar quasidimers was included in the calculations. Rate constants in agreement with experiment within a factor of 2 were obtained. It was noted that it would have been possible to obtain the negative temperature coefficient of the rate constant, in agreement with experiment, by choosing the proper ϵ_{Br-Ar} .

Most recent calculations^{19f} made use of preliminary information from molecular beam data to better define the potential energy surface of the recombination process. By using the classical 3-D trajectory method with an improved sampling technique, reasonable agreement between computed and experimental rate constants was obtained for the Br recombination in He, Ar and Xe. The following L-J parameters were used for the Br-RG potentials: RG=He, $\epsilon=0.15$ kcal/mole, $\sigma=2.88$ Å; RG=Ar, $\epsilon=0.6$ kcal/mole, $\sigma=3.32$ Å; RG=Xe, $\epsilon=0.7$ kcal/mole, $\sigma=3.62$ Å.

It should be noted that in all the theoretical recombination studies, the halogen-rare gas interaction has always been described in terms of a single potential energy curve. However, the effect of multiple potential curves or surfaces in systems containing open shell atoms needs careful consideration.⁵²

V. CONCLUDING REMARKS

The experiments reported in this paper represent a further step towards the characterization of the ground state RG-X potential energy curves. Information unattainable by other techniques has been provided. A comparison between theory and experiment shows the existence of discrepancies in the predicted properties of both the excited and ground states of these molecules. In light of the many factors which may determine the peculiar RG-X electronic structure and which are not easily handled by the current theoretical treatment, these discrepancies are not surprising. An accurate knowledge of the ground state potential manifold, as derived from crossed molecular beam experiments, may serve the useful purpose of furnishing a reference for refining the excited state potentials in the theoretical simulation of emission spectra. It should be noted that there exists quite a good agreement between the slopes of the ground state potential estimated from diffuse spectra simulation and those derived from the experiments described in this paper. What our study adds is the absolute scale of the internuclear distance. Thus, the minimum positions and shapes of the excited states can be more reliably derived and more useful comparison made with the ground state alkali halide electronic structure and theoretical predictions. A fruitful "coupling" with spectroscopic techniques⁵³ and theoretical calculations should contribute to a more complete understanding of the RG-X electronic structure.

The investigation of the I-RG system is now underway,⁵⁰ which will complete the matrix of our scattering studies on the rare gas-halogen diatomic molecules. From the analysis of rows and columns of this matrix, the effect of the halogen and/or rare gas will be more fully discussed and the subtle transitions from van der Waals forces to chemical forces more clearly examined.

ACKNOWLEDGMENTS

P. C. wishes to thank Dr. Christopher H. Becker for many enjoyable and useful discussions. This work was supported by the Office of Naval Research (Contract No. N00014-77-C-0101), and by the Director, Office of Energy Research, Office of Basic Energy Sciences, Chemical Sciences Division of the U.S. Department of Energy under Contract Number W-7405-ENG-48. P.C. acknowledges a fellowship from the Italian Ministry of Education and travel support from NATO Grant No. 1444.

REFERENCES AND FOOTNOTES

1. For emission see e.g., (a) L. A. Kuznetsova, Y. Y. Kuzyakov, V. A. Shpanskii, and V. M. Khutoretskii, *Vestn. Mosk. Univ. Ser. II Khim.* 19, 19 (1964); (b) M. F. Golde and B. A. Thrush, *Chem. Phys. Lett.* 29, 486 (1974); (c) J. E. Velazco and D. W. Setser, *J. Chem. Phys.* 62, 1990 (1975); (d) J. J. Ewing and C. A. Brau, *Phys. Rev. A* 12, 129 (1975); (e) J. Tellinghuisen, A. K. Hays, J. M. Hoffman, and G. C. Tisone, *J. Chem. Phys.* 65, 4473 (1976).
2. For laser action, see e.g., (a) S. K. Searles and G. A. Hart, *Appl. Phys. Lett.* 27, 243 (1975); (b) J. J. Ewing and C. A. Brau, *Appl. Phys. Lett.* 27, 350 (1975); (c) J. A. Mangano and J. H. Jacob, *Appl. Phys. Lett.* 27, 495 (1975); (d) G. C. Tisone, A. K. Hays, and J. M. Hoffman, *Opt. Commun.* 15, 188 (1975); (e) C. A. Brau and J. J. Ewing, *Appl. Phys. Lett.* 27, 435 (1975); (f) J. M. Hoffman, A. K. Hays, and G. C. Tisone, *Appl. Phys. Lett.* 28, 538 (1976); (g) J. R. Murray and H. T. Powell, *Appl. Phys. Lett.* 29, 252 (1976); (h) J. G. Eden and S. K. Searles, *Appl. Phys. Lett.* 29, 350 (1976).
3. (a) T. H. Dunning and P. J. Hay, *J. Chem. Phys.* 69, 134 (1978);
(b) P. J. Hay and T. H. Dunning, *J. Chem. Phys.* 69, 2209 (1978);
(c) M. Krauss, *J. Chem. Phys.* 29, 350 (1976).
4. M. F. Golde, *J. Mol. Spectr.* 58, 261 (1975).
5. C. A. Brau and J. J. Ewing, *J. Chem. Phys.* 63, 4640 (1975).
6. J. Tellinghuisen, P. C. Tellinghuisen, J. A. Coxon, J. E. Velazco, and D. W. Setser, *J. Chem. Phys.* 68, 5187 (1978).
7. A. L. Smith and P. C. Kobrinsky, *J. Mol. Spectr.* 69, 1 (1978).

8. C. H. Becker, P. Casavecchia, and Y. T. Lee, *J. Chem. Phys.* 69, 2377 (1978).
9. A. Sur, A. K. Hui, and J. Tellinghuisen, *J. Mol. Spectr.* 74, 465 (1979). J. Tellinghuisen, J. M. Hoffman, G. C. Tisone, and A. K. Hays, *J. Chem. Phys.* 64, 2484 (1976).
10. C. H. Becker, J. J. Valentini, P. Casavecchia, S. J. Sibener and Y. T. Lee, *Chem. Phys. Lett.* 61, 1 (1979).
11. See, for example: P. J. Hay, W. R. Wadt, and T. H. Dunning, *Ann. Rev. Phys. Chem.* 30, 311 (1979), and references cited therein.
12. See: J. Tellinghuisen and M. R. McKeever, *Chem. Phys. Lett.* 72, 94 (1980), and references cited therein.
13. W. Kutzelnigg, *Faraday Disc. Chem. Soc.* 62, 185 (1977).
14. R. B. Bernstein and J. T. Muckerman, *Adv. Chem. Phys.* 12, 389 (1967); Ch. Schlier, *Ann. Rev. Phys. Chem.* 20, 191 (1969); U. Buck, *Adv. Chem. Phys.* 30, 313 (1975); H. Pauly, in Atom-Molecule Collision Theory, Ed. R. B. Bernstein (Plenum Press, New York, 1979), Chapter 4.
15. C. H. Becker, P. Casavecchia and Y. T. Lee, *J. Chem. Phys.* 70, 2986 (1979).
16. M. P. Casassa, M. F. Golde and A. Kvaran, *Chem. Phys. Lett.* 59, 51 (1978).
17. M. F. Golde and A. Kvaran, *J. Chem. Phys.* 72, 434 (1980); *J. Chem. Phys.* 72, 442 (1980).
18. J. Tellinghuisen, A. K. Hays, J. M. Hoffman and G. C. Tisone, *J. Chem. Phys.* 65, 4473 (1976).

19. (a) A. G. Clarke and G. Burns, J. Chem. Phys. 55, 4717 (1971);
(b) A. G. Clarke and G. Burns, J. Chem. Phys. 56, 4636 (1972);
(c) V. H. Shui, J. P. Appleton, and J. C. Keck, Proc. 13th Symp. Combust. 1970 (1971), p. 21; (d) A. G. Clarke and G. Burns, J. Chem. Phys. 58, 1908 (1973); (e) W. H. Wong and G. Burns, J. Chem. Phys. 59, 2974 (1973); (f) D. T. Chang and G. Burns, Can. J. Chem. 54, 1535 (1976).
20. For some different viewpoints on theories of termolecular recombination, see also, G. Porter and J. A. Smith, Nature 184, 446 (1959); J. C. Keck, Adv. Chem. Phys. 13, 85 (1967); D. L. Bunker, Theory of Elementary Gas Reaction Rates (Pergamon, Oxford, 1966), Chapter 4; W. H. Wong and G. Burns, J. Chem. Phys. 58, 4459 (1973); and E. E. Nikitin, Theory of Elementary Atomic and Molecular Processes in Gases (Clarendon Press, Oxford, 1974), Chapter 7.
21. R. K. Boyd and G. Burns, J. Phys. Chem. 83, 88 (1979).
22. R. K. Sparks, Ph.D. Thesis, University of California, Berkeley, California, 1979.
23. Y. T. Lee, J. D. McDonald, P. R. Le Breton and D. R. Herschbach, Rev. Sci. Instr. 40, 1402 (1969).
24. J. J. Valentini, M. J. Coggiola, and Y. T. Lee, Rev. Sci. Instr. 48, 58 (1977).
25. C. H. Becker, P. Casavecchia, Y. T. Lee, R. E Olson, and W. A. Lester, Jr., J. Chem. Phys. 70, 5477 (1979).
26. See, e.g., E.H.S. Burhop, in Quantum Theory, D. R. Bates, editor (Academic Press, New York, 1961), Vol. I, Ch. 9.

27. V. Aquilanti and G. Grossi, J. Chem. Phys. 73, 1165 (1980).
28. V. Aquilanti, P. Casavecchia, G. Grossi and A. Lagana, J. Chem. Phys. 73, 1173 (1980).
29. V. Aquilanti, E. Luzzatti, F. Pirani, and G. G. Volpi, J. Chem. Phys. 73, 1181 (1980).
30. R.H.G. Reid, J. Phys. B 8, L 493 (1975); R.H.G. Reid and R. F. Franklin, ibid. 11, 55 (1978); C. Harel, V. Lopez, R. McCarroll, A. Riera, and P. Wahnon, ibid. 11, 71 (1978); B. Amace and C. Bottcher, ibid. 11, 1249 (1978).
31. J. C. Slater and J. G. Kirkwood, Phys. Rev. 37, 682 (1931); see also K. S. Pitzer, Adv. Chem. Phys. 2, 59 (1959).
32. R. R. Teachout and R. T. Pack, At. Data 3, 195 (1971).
33. J. S. Cohen and R T. Pack, J. Chem. Phys. 61, 2372 (1974).
34. H. J. Werner and W. Meyer, Phys. Rev. A13, 13 (1976).
35. E. A. Reinsch and W. Meyer, Phys. Rev. A14, 915 (1976).
36. R. K. Preston, C. Sloane, and W. H. Miller, J. Chem. Phys. 60, 4961 (1974).
37. C. Froese Fisher, At. Data 4, 301 (1972). The quadrupole moment is obtained in terms of the mean value of $\langle r^2 \rangle$ over the valence orbital.
38. (a) C. H. Becker, R. J. Buss and Y. T. Lee, Materials and Molecular Research Division, Lawrence Berkeley Laboratory, University of California, Annual Report 1977 (LBL-7355), pp. 384-386; (b) See also, R. A. Aziz, J. Presley, U. Buck and J. Schleusener, J. Chem. Phys. 70, 4737 (1979).

39. J. A. Barker, R. O. Watts, J. K. Lee, T. P. Schafer and Y. T. Lee, *J. Chem. Phys.* 61, 3081 (1974).
40. C. H. Becker, R. J. Buss and Y. T. Lee, to be published.
41. See also, J. M. Farrar, T. P. Schafer and Y. T. Lee, *AIP Conference Proceedings*, No. 11, *Transport Phenomena* (1973), Ed. J. Kestin, p. 279.
42. R. J. Le Roy, Molecular Spectroscopy (The Chemical Society, London, 1973), Vol. I, pp. 113-176.
43. P. Brumer and M. Karplus, *J. Chem. Phys.* 58, 3903 (1973).
44. D. J. Ehrlich and R. M Osgood, Jr., *J. Chem. Phys.* 73, 3038 (1980).
45. M. Krauss and B. Liu, *Chem. Phys. Lett.* 44, 257 (1976).
46. B. Liu and A. D. McLean, *J. Chem. Phys.* 59, 4557 (1973).
47. Present work.
48. (a) D. Klingler, H. H. Nakano, D. L. Huestis, W. K. Bischel, R. M. Hill and C. K. Rhodes, *Appl. Phys. Lett.* 33, 39 (1978); (b) J. H. Kolts and D. W. Setzer, *J. Phys. Chem.* 82, 1776 (1978); (c) H. C. Brashears and D. W. Setzer, *Appl. Phys. Lett.* 33, 821 (1978); (d) R. W. Waynant and J. G. Eden, *IEEE J. Quantum Electron.* 15, 61 (1979); (e) H. C. Brashears, Jr. and D. W. Setzer, *J. Phys. Chem.* 84, 224 (1980).
49. B. S. Ault and L. Andrews, *J. Chem. Phys.* 65, 4192 (1976).
50. P. Casavecchia, H. Guozhong, R. K. Sparks and Y. T. Lee, to be published.
51. For an up to date discussion on the current status of experiment and theory, see Ref. 21 and references cited therein.

52. D. G. Truhlar, J. Chem. Phys. 56, 3189 (1972); J. T. Muckerman and M. D. Newton, J. Chem. Phys. 56, 3191 (1972).
53. Also the laser induced fluorescence technique in supersonic free jets can yield useful information on the potential energy curves of the RG-X. See, for example: D. L. Monts, L. M. Ziurys, S. M. Beck, M. G. Liverman, and R. E. Smalley, J. Chem. Phys. 71, 4057 (1979).

Table I. Beam Characteristics and Center-of-Mass Collision Energies

Beam Br	Effective Temperature (K)	Stagnation Pressure (Torr)	Mach Number	Peak Velocity (10 ⁴ cm/s)	$\Delta V/V$ (FWHM)	Average Collision Energies E (Kcal/mole)
				(FWHM)	Br-Xe Br-Kr Br-Ar	
9% Br ₂ + 91% Xe	1206	700	8.1	7.86	0.17	4.2
7.7% Br ₂ + 92.3% Ar	3017	820	10.5	12.42	0.15	9.8
6.3% Br ₂ + 93.7% He	8505	1000	9.1	20.86	0.17	26.4
					-	14.8
Xe	298	200	17.3	3.08	0.10	
Kr	296	300	17.6	3.81	0.10	
Ar	296	300	17.3	5.55	0.10	

Table II. Interaction Potential Parameters for Br-Ar, Kr, Xe, and Kr-Ar, Kr, Xe

System	Br-Ar			Br-Kr			Br-Xe			Kr-Ar			Kr-Kr			Kr-Xe		
	1 ₂ ¹	1 ₂ ³	x ₂ ¹	1 ₂ ⁺														
Electronic State	x ₂ ¹	1 ₂ ³	x ₂ ¹	1 ₂ ⁺														
a (kcal/mole)	0.380	0.260	0.460	0.345	0.645	0.425	0.338	0.394	0.401	c	d	0.465						
r _m (Å)	3.73	3.93	3.90	4.10	3.80	4.11	3.83	4.11	4.01			4.12						
b ₁	7.10	7.90	5.70	7.20	4.30	6.65	6.4					6.0						
b ₂	7.30	7.50	7.70	7.80	6.50	6.80	6.78					6.3						
x ₁	1.0255	1.0166	1.0231	1.0217	1.0405	1.0284	1.0118					1.0289						
x ₂	1.700	1.700	1.700	1.600	2.000	1.900	1.45					1.37						
c ₆ (kcal/mole·Å ⁶)	1504.	1616.	2130.	2294.	3223.	3476.	1304.					2800.						
c ₈ (kcal/mole·Å ⁸)	8730.	8730.	13050.	13050.	21570.	21570.	7293.					15000.						

a Ref. 38.

b Ref. 41.

c Ref. 39. This Kr-Kr potential should be more accurate than that of Ref. 41. Only c and r_m are reported since the potential is not of the MSV form. The repulsive walls are very similar.

d Ref. 40.

FIGURE CAPTIONS

Figure 1. Velocity distributions of bromine atom beams used on the experiments. Solid circles are from TOF data and lines are results of parametric fits to the deconvoluted distributions for the Mach numbers and temperatures given in Table I. Peaks are normalized to unity.

Figure 2. Laboratory angular distributions of scattered Br for the $\text{Br}({}^2\text{P}) + \text{Xe}({}^1\text{S})$ system at three collision energies. Solid circles are data points and the solid curves are calculated from the best fit potentials of Table II, averaging over angular and velocity distributions of experimental conditions. Dashed and dashed-dotted curves represent the relative contribution to $I(\text{H})$ of the $\text{X} 1/2$ and $\text{I} 3/2$ potentials of Table II, according to Eq. (1).

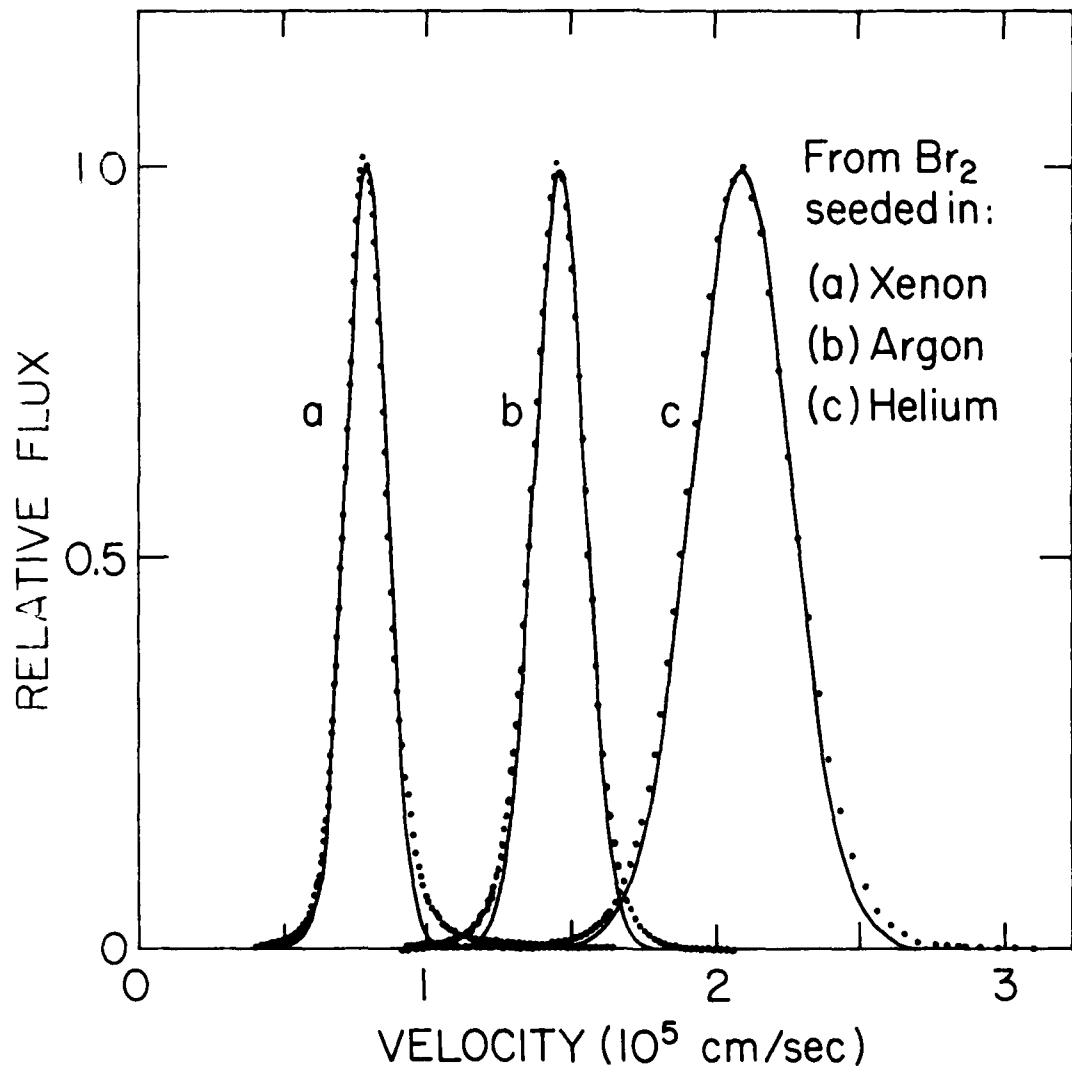
Figure 3. Laboratory angular distributions of scattered Br for the $\text{Br}({}^2\text{P}) + \text{Kr}({}^1\text{S})$ system at two collision energies. Symbols are the same as in Fig. 2.

Figure 4. Laboratory angular distributions of scattered Br for the $\text{Br}({}^2\text{P}) + \text{Ar}({}^1\text{S})$ system at two collision energies. Symbols are the same as in Fig. 2.

Figure 5. Interaction potentials for $\text{Br}({}^2\text{P}_{3/2}) + \text{Xe}({}^1\text{S}_0)$ obtained from experimental results shown in Fig. 2. Note the scale change at $V(r)$ higher than 0.1 kcal/mole.

Figure 6. Interaction potentials for $\text{Br}({}^2\text{P}_{3/2}) + \text{Kr}({}^1\text{S}_0)$ obtained from experimental results shown in Fig. 3.

Figure 7. Interaction potentials for $\text{Br}({}^2\text{P}_{3/2}) + \text{Ar}({}^1\text{S}_0)$ obtained from experimental results shown in Fig. 4.



XBL 813-8456

Fig. 1

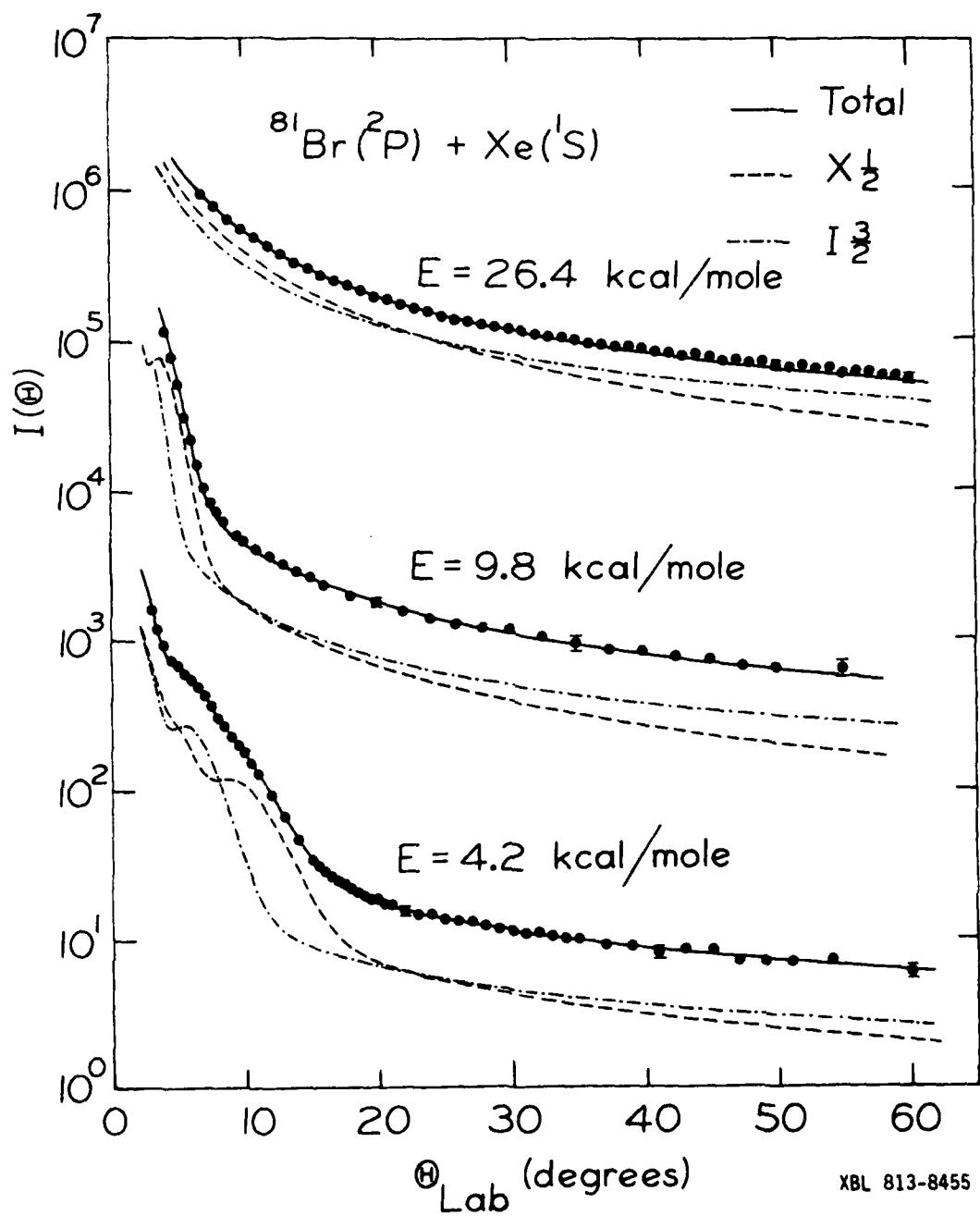


Fig. 2

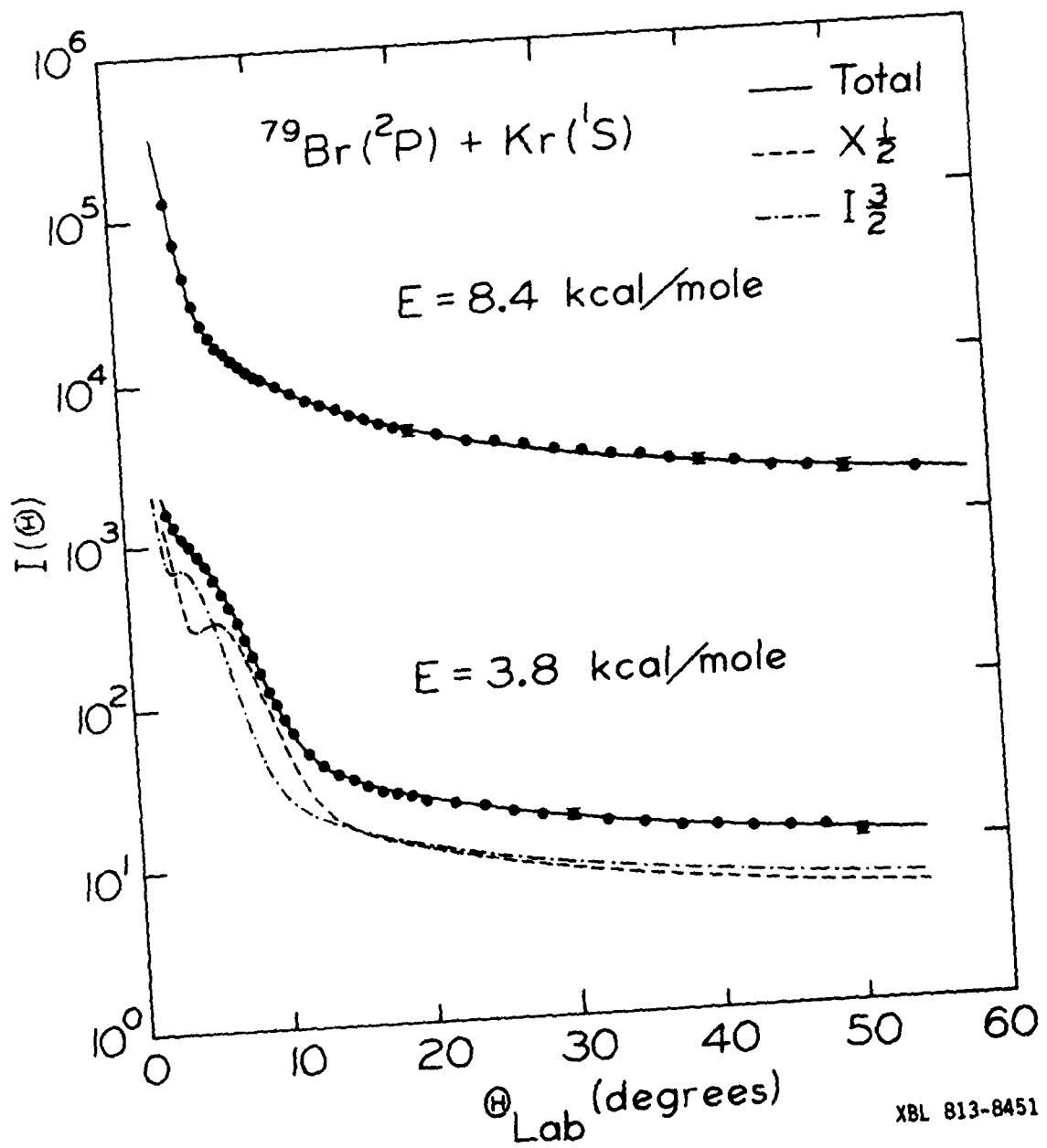


Fig. 3

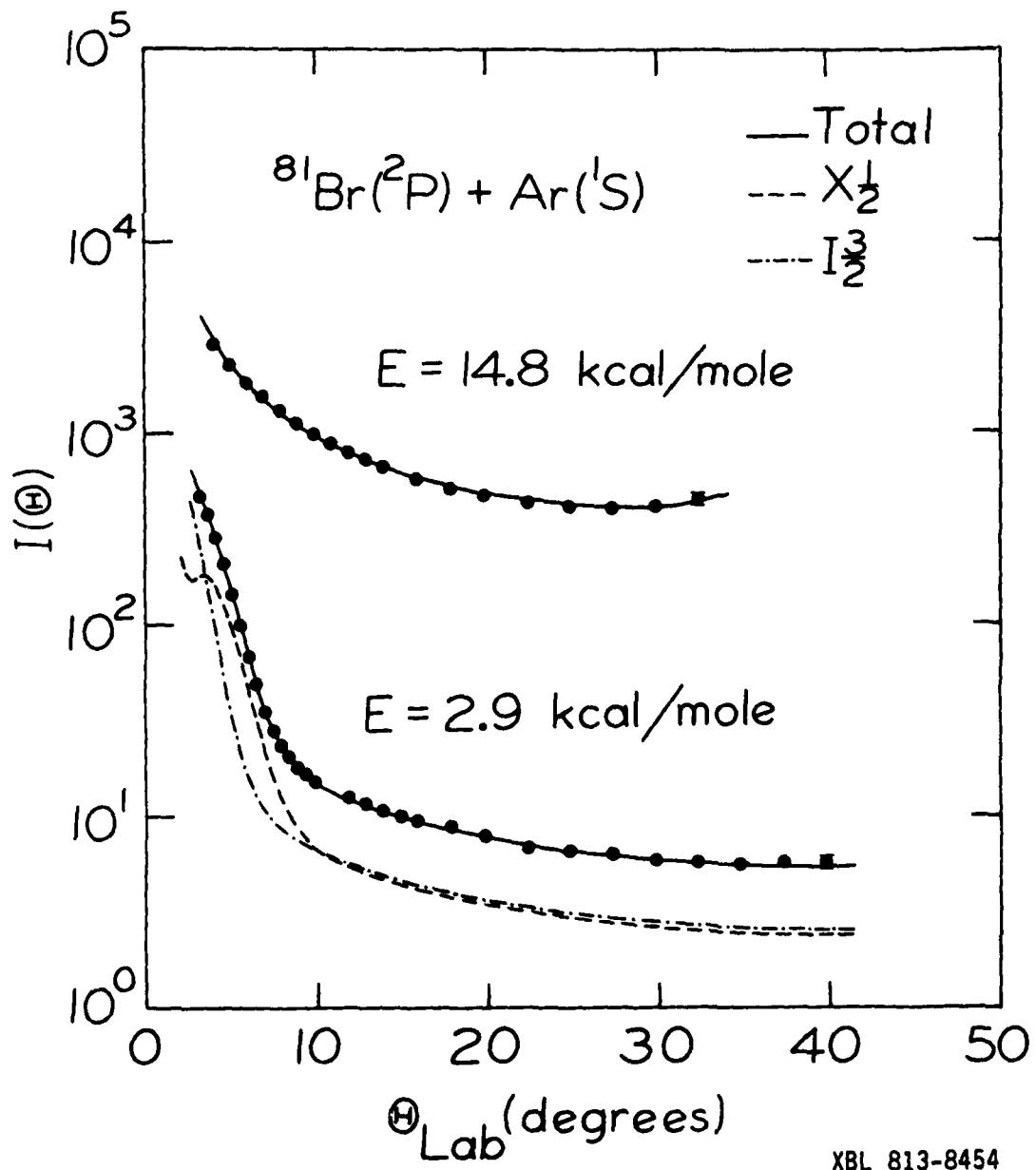


Fig. 4

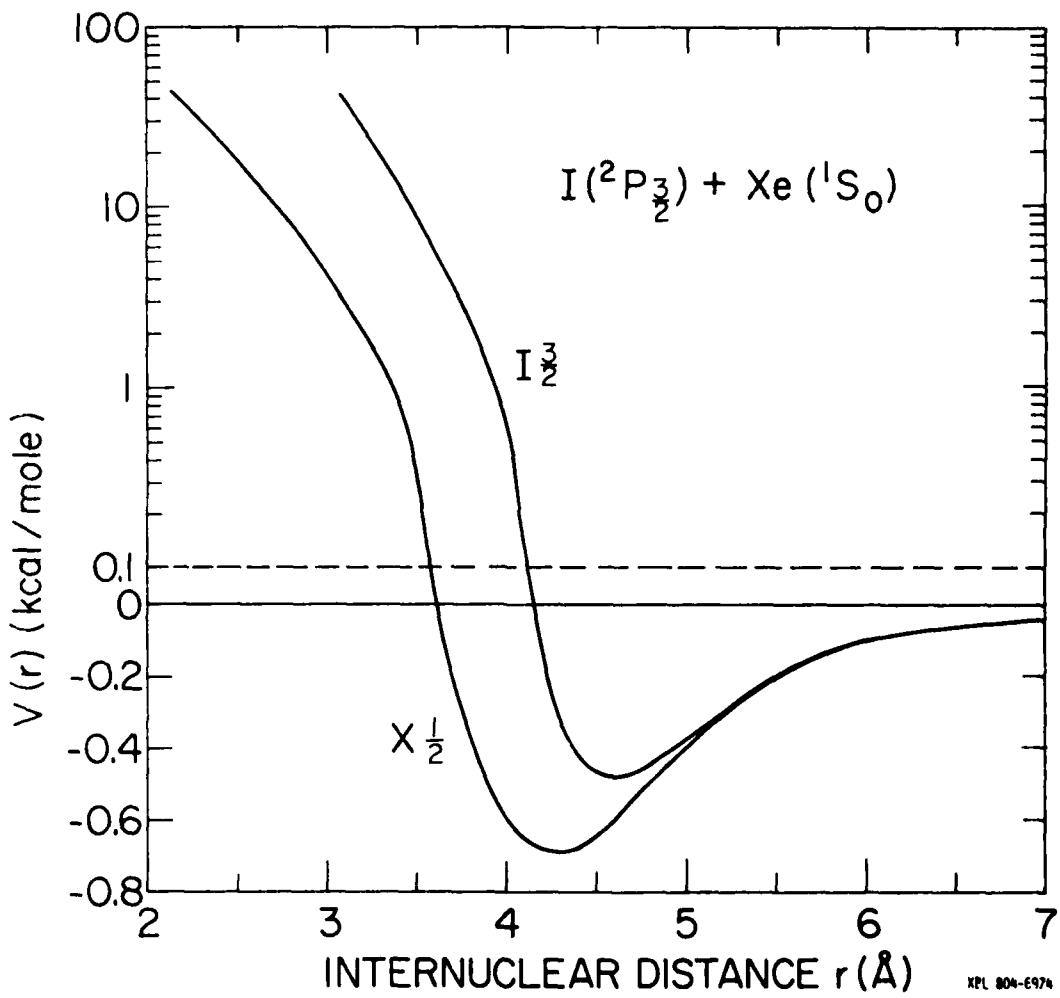


Fig. 5

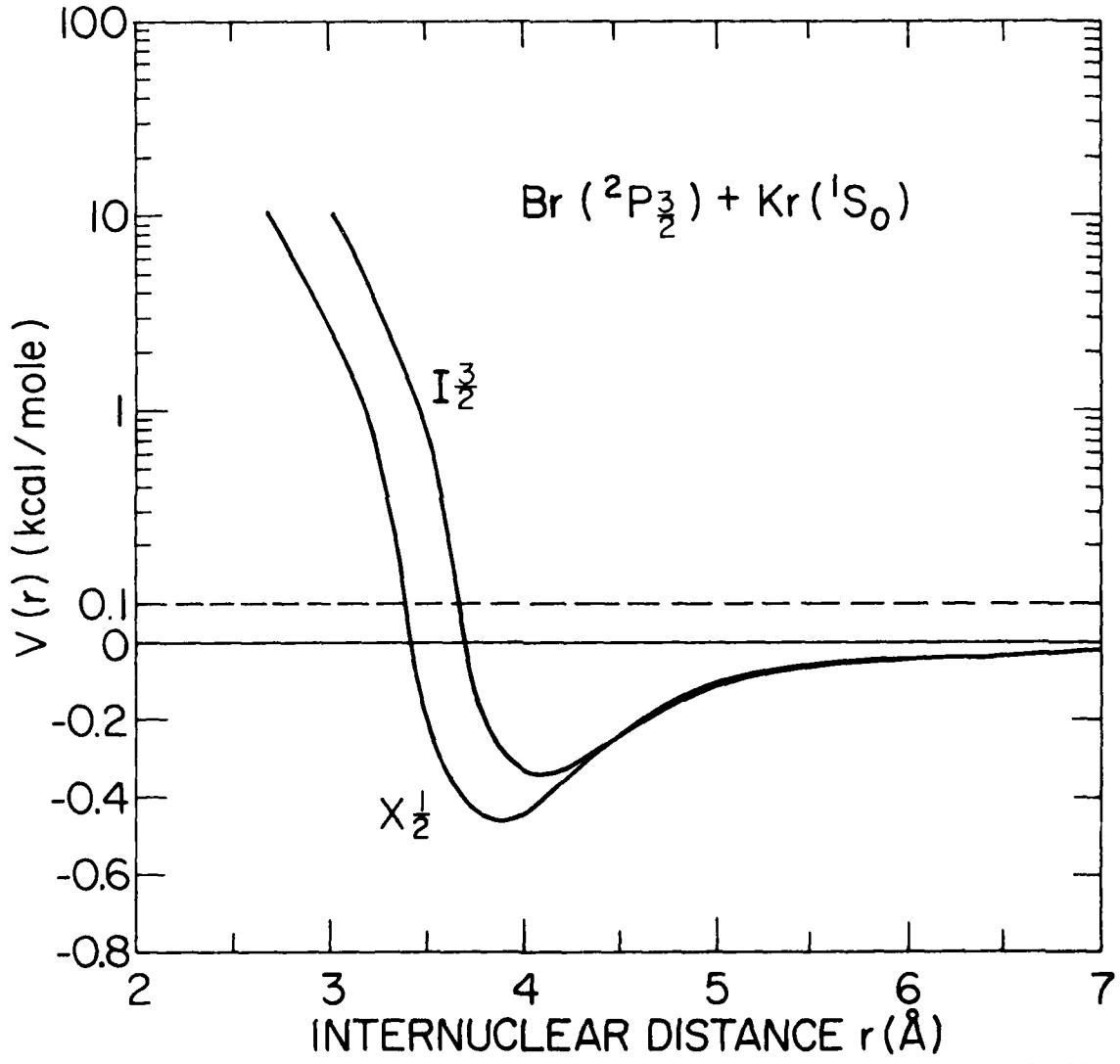


Fig. 6

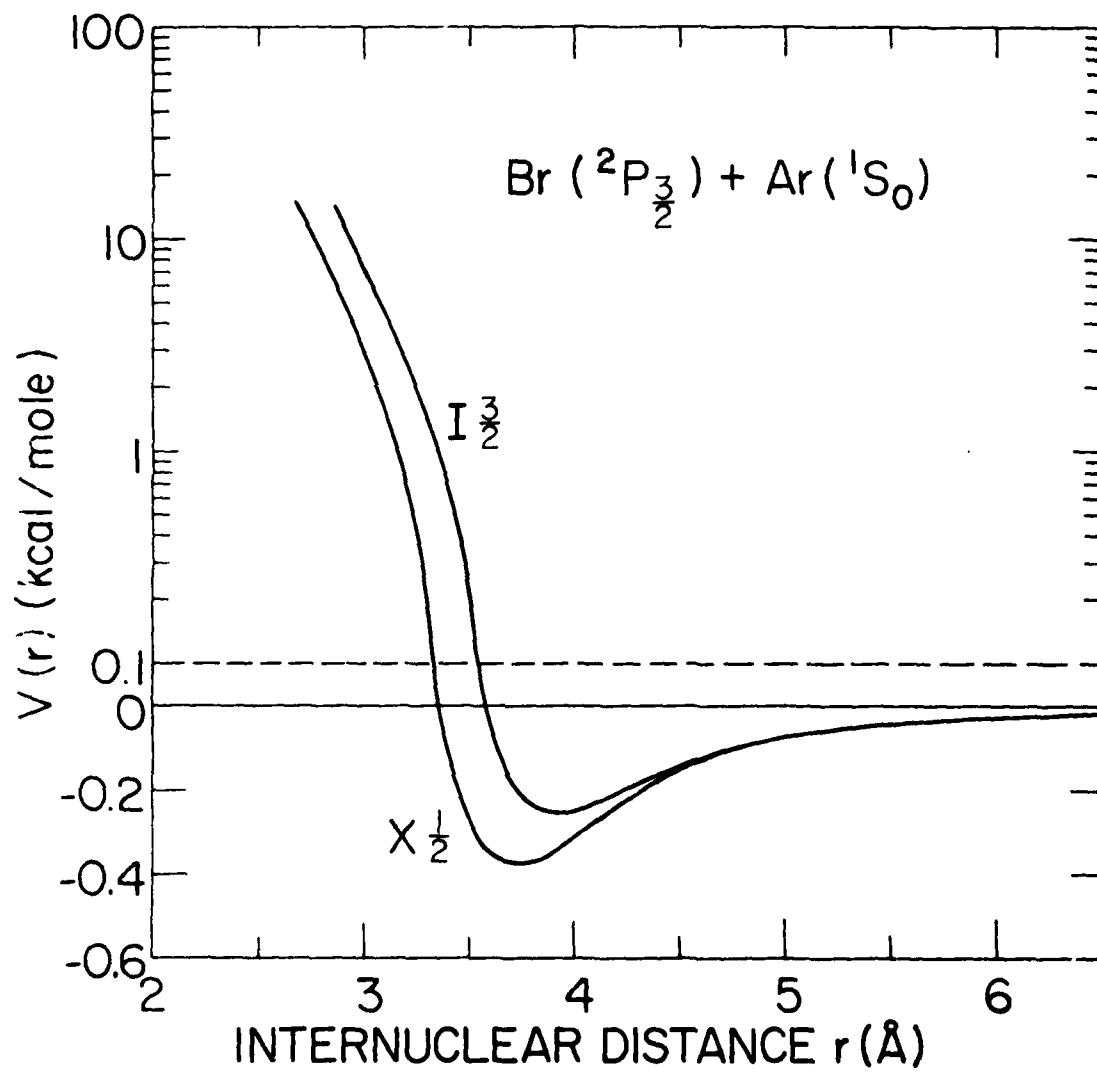


Fig. 7

March 1976

Distribution List for ARPA Order 2840 Unclassified Contracts

Office of Naval Research Department of the Navy Attn: Physics Program Arlington, Virginia 22217	3 copies
Office of the Director of Defense Research and Engineering Information Office Library Branch The Pentagon Washington, D. C. 20301	1 copy
U. S. Army Research Office Box CM Duke Station Durham, N. C. 97706	1 copy
Defense Documentation Center Cameron Station Alexandria, VA 22314	12 copies
Defender Information Analysis Center Battelle Memorial Institute 505 King Avenue Columbus, OH 43201	1 copy
Director Office of Naval Research Branch Office 536 South Clark Street Chicago, IL 60615	1 copy
San Francisco Area Office Office of Naval Research 760 Market Street, Room 447 San Francisco, CA 94102	1 copy
Air Force Office of Scientific Research Department of the Air Force Washington, D. C. 22209	1 copy
Office of Naval Research Branch Office Attn: Dr. Robert Behringer 1030 East Green Street Pasadena, CA 91106	1 copy
Code 102 1P(ONRL) Office of Naval Research 800 North Quincy Street Arlington, VA 22217	6 copies

Distribution List for ARPA Order 2840 Unclassified Contracts

Defense Advanced Research Projects Agency
1400 Wilson Blvd.
Arlington, VA 22209
Attn: Director, Laser Division 1 copy

ODDR&E
The Pentagon
Washington, D. C. 20301
Attn: (Ass't. Dir. (Space and Advanced Systems) 1 copy

Office of the Ass't Secretary of Defense
System Analysis (Strategic Programs)
Washington, D. C. 20301
Attn: Mr. Gerald R. McNichols 1 copy

U. S. Arms Control and Disarmament Agency
Dept. of State Bldg., RM 4931
Washington, D. C. 20451
Attn: Dr. Charles Henkin 1 copy

Energy Research Development
Division of Military Applications
Washington, D. C. 20545 1 copy

National Aeronautics and Space Administration
Lewis Research Center
Cleveland, OH 44135
Attn: Dr. John W. Dunning, Jr.
(Aerospace Res. Engineer) 1 copy

National Aeronautics & Space Administration
Code RR, FOB 10B
600 Independence Ave., S. W.
Washington, D. C. 20546 1 copy

National Aeronautics & Space Administration
Ames Research Center
Moffet Field, CA 94035
Attn: Mr. Robert L. McKenzie
Dr. Kenneth W. Billman 1 copy
1 copy

Department of the Army
Office of the Chief of RD&A
Washington, D. C. 20310
Attn: DARD-DD
DAMA-WSM-T 1 copy
1 copy

Distribution List for ARPA Order 2840 Unclassified Contracts

Department of the Navy Office of the Chief of Naval Operations Pentagon 5C739 Washington, D. C. 20350 Attn: (OP 982F3)	1 copy
Office of Naval Research Attn: Dr. Fred Quelle 495 Summer Street Boston, MA 02210	1 copy
Department of the Navy Deputy Chief of Naval Material (Dev.) Washington, D. C. 20360 Attn: Mr. R. Gaylord (MAT 032B)	1 copy
Naval Missile Center Point Mugu, CA 93042 Attn: Gary Gibbs (Code 0146)	1 copy
Naval Research Laboratory Washington, D. C. 20375 Attn: (Code 5503-LTP0) Dr. P. Livingston (Code 5560) Dr. A. I. Schindler (Code 6330) Dr. H. Shenker (Code 6530) Mr. D. J. McLaughlin (Code 5560) Dr. John L. Walsh (Code 5503) Technical Library	1 copy 1 copy 1 copy 1 copy 1 copy 1 copy 1 copy
Commander, Naval Sea Systems Command Department of the Navy Washington, D. C. 20362 Attn: CAPT A. Skolnick, USN (PM-22)	1 copy
Superintendent Naval Postgraduate School Monterey, CA 93940 Attn: Library (Code 2124)	1 copy
Naval Surface Weapons Center White Oak Silver Spring, MD 20910 Attn: Dr. Leon H. Schindel (Code 310) Dr. E. Leroy Harris (Code 313) Mr. K. Enkenhus (Code 034) Mr. J. Wise (Code 047) Technical Library	1 copy 1 copy 1 copy 1 copy 1 copy

Distribution List for ARPA Order 2840 Unclassified Contracts

U. S. Naval Weapons Center China Lake, CA 93555 Attn: (Code 5114) Technical Library	1 copy 1 copy
HQ USAF (AF/RDPS) The Pentagon Washington, D. C. 20330 Attn: Lt. Col. A. J. Chiota	1 copy
HQ AFSC/XRLW Andrews AFB Washington, D. C. 20331 Attn: MAJ J. M. Walton	1 copy
HQ AFSC (DLCAW) Andrews AFB Washington, D. C. 20334 Attn: CAPT Pulley	1 copy
Air Force Weapons Lab Kirtland AFB, NM 87117 Attn: LR AL	4 copies 2 copies
HQ SAMSO P. O. Box 92960, Worldway Postal Center Los Angeles, CA 90009 Attn: CAPT T. J. Ernst (DYD)	1 copy
AF Avionics Lab (TEO) Wright Patterson AFB, OH 45433 Attn: Mr. K. Hutchinson	1 copy
Dept. of the Air Force Air Force Materials Lab. (AFSC) Wright Patterson AFB, OH 45433 Attn: MAJ Paul Elder (LPS) Laser Window Group	1 copy
HQ Aeronautical Systems Div. Wright Patterson AFB, OH 45433 Attn: XRF - Mr. Clifford Fawcett	1 copy
Rome Air Development Command Griffiss AFB Rome, N. Y. 13441 Attn: Mr. R. Urtz (OCSE)	1 copy

Distribution List for ARPA Order 2840 Unclassified Contracts

HQ Electronics Systems Div. (ESD)

L. G. Hanscom Field

Bedford, MA 01731

Attn: CAPT A. R. Tobin (XRE)

Technical Library

1 copy

1 copy

Air Force Rocket Propulsion Lab

Edwards AFB, CA 93523

Attn: B. R. Bornhorst, (LMCG)

1 copy

Air Force Aero Propulsion Lab.

Wright Patterson AFB, OH 45433

Attn: COL Walter Moe (CC)

1 copy

Dept. of the Air Force

Foreign Technology Division

Wright Patterson AFB, OH 45433

Attn: PDM

1 copy

CINCSAC/INEP

Offutt AFB, NE 68113

1 copy

Commandant of the Marine Corps

Scientific Advisor (Code RD-1)

Washington, D. C. 20380

1 copy

USAF/INA
KA

Washington, D. C. 20330

Attn: LT COL W. M. Truesdell

1 copy

Defense Intelligence Agency

Washington, D. C. 20301

Attn: Mr. Seymour Berler (DTIB)

1 copy

Central Intelligence Agency

Washington, D. C. 20505

Attn: Mr. Julian C. Nall

1 copy

Aerodyne Research Inc.

Northwest Industrial Park

Burlington, MA 01803

Attn: C. E. Kolb

1 copy

Analytic Services, Inc.

5613 Leesburg Pike

Falls Church, VA 22041

Attn: Mr. John Davis

1 copy

Distribution List for ARPA Order 2840 Unclassified Contracts

Aerospace Corp. P.O. Box 92957 Los Angeles, CA 90009 Attn: Dr. G. P. Millburn	1 copy
Airesearch Manuf. Co. 9851 - 9951 Sepulveda Blvd. Los Angeles, CA 90009 Attn: Mr. A. Colin Stancliffe	1 copy
Atlantic Research Corp. Shirley Highway at Edsall Road Alexandria, VA 22314 Attn: Mr. Robert Naismith	1 copy
AVCO Everett Research Lab. 2385 Revere Beach Parkway Everett, MA 02149 Attn: Dr. George Sutton Dr. Jack Daugherty Dr. D. Hamilton Dr. D. Trainor	1 copy 1 copy 1 copy 1 copy
Battelle Columbus Laboratories 505 King Avenue Columbus, OH 43201 Attn: Mr. Fred Tietzel (STOILAC)	1 copy
Bell Aerospace Co. Buffalo, N.Y. 14240 Attn: Dr. Wayne C. Solomon	1 copy
Boeing Co. P.O. Box 3999 Seattle, WA 98124 Attn: Mr. M. I. Gamble (2-1460, MS 8C-88)	2 copies
Electro-Optical Systems 300 North Halstead Pasadena, CA 91107 Attn: Dr. Andrew Jensen	1 copy
Environmental Research Institute of Michigan P.O. Box 618 Ann Arbor, MI 48107 Attn: IRIA (Laser Library)	1 copy

Distribution List for ARPA Order 2840 Unclassified Contracts

General Electric Co.
Space Division
P.O. Box 8555
Philadelphia, PA 19101
Attn: Dr. R. R. Sigismonti 1 copy
 Dr. C. E. Anderson 1 copy

General Electric Co.
100 Plastics Avenue
Pittsfield, MA 01201
Attn: Mr. D. G. Harrington
(Room 1044) 1 copy

General Research Corp.
P.O. Box 3587
Santa Barbara, CA 93105
Attn: Dr. R. Holbrook

General Research Corp.
1501 Wilson Blvd., Suite 700
Arlington, Va. 22209
Attn: Dr. Giles F. Crimi 1 copy

Hercules, Inc.
Industrial System Dept.
Wilmington, DE 19899
Attn: Dr. R. S. Voris

Hercules, Inc.
P. O. Box 210
Cumberland, MD 21502
Attn: Dr. Ralph R. Preckel

Hughes Research Labs.
3011 Malibu Canyon Road
Malibu, CA 90265

Attn: Dr. D. Forster 1 copy
Dr. A. N. Chester 1 copy
Dr. Viktor Evtuhov 1 copy

Hughes Aircraft Co.
Centinela and Teale Streets
Bldg. 6 MS E-125
Culver City, CA 90230
Attn: Dr. J. A. Alcalay

Distribution List for ARPA Order 2840 Unclassified Contracts

Institute for Defense Analysis 400 Army Navy Drive Arlington, VA 22202 Attn: Dr. Alvin Schnitzler	1 copy
Johns Hopkins University Applied Physics Lab 8621 Georgia Ave. Silver Spring, MD 20910 Attn: Dr. Albert M. Stone Dr. R. E. Gorozdos	1 copy 1 copy
Lawrence Livermore Lab. P.O. Box 808 Livermore, CA 94550 Attn: Dr. R. E. Kidder Dr. E. Teller Dr. Joe Fleck Dr. John Emmett Dr. William F. Krupke	1 copy 1 copy 1 copy 1 copy 1 copy
Los Alamos Scientific Lab P. O. Box 1663 Los Alamos, NM 87544 Attn: Dr. Keith Boyer	1 copy
Iulejian and Associates, Inc. Del Amo Financial Center Suite 500 21515 Hawthorne Blvd. Torrance, CA 90503	1 copy
Lockheed Palo Alto Res. Lab. 3251 Hanover Street Palo Alto, CA 94303 Attn: L. R. Lunsford Orgn. 52-24 Bldg. 201	1 copy
Mathematical Sciences Northwest, Inc. P. O. Box 1887 Bellevue, WN 98009 Attn: Mr. Peter H. Rose Dr. Abraham Hertzberg	1 copy 1 copy

Distribution List for ARPA Order 2840 Unclassified Contracts

Martin Marietta Corp.
P.O. Box 179
Mail Station 0471
Denver, CO 80201
Attn: Mr. Stewart Chapin
Mr. Scott Giles

1 copy
1 copy

Massachusetts Institute of Technology
LINCOLN LAB
P.O. Box 73
Lexington, MA 02173
Attn: Dr. L. C. Marquet
Dr. J. Freedman
Dr. G. P. Dinneen
Dr. R. H. Rediker

1 copy
1 copy
1 copy
1 copy

Mc Donnell Douglas Astronautics Co.
5301 Bolsa Avenue
Huntington Beach, CA 92647
Attn: Mr. P. L. Klevatt
Dept. A3-830-BNFO, M/S 9

1 copy

Mc Donnell Douglas Research Labs.
Dept. 220, Box 516
St. Louis, MO 63166
Attn: Dr. D. P. Ames

1 copy

MITRE Corp
P.O. Box 208
Bedford, MA 01730
Attn: Mr. A. C. Cron

1 copy

Northrop Corp.
3401 West Broadway
Hawthorne, CA 90250
Attn: Dr. Gerard Hasserjian
Laser Systems Dept.

1 copy

Physical Sciences Inc.
18 Lakeside Office Park
Wakefield, MA 01880
Attn: Dr. Anthony N. Pirri

1 copy

RAND Corp
1700 Main Street
Santa Monica, CA 90406
Attn: Dr. Claude R. Culp/Mr. G. A. Carter

1 copy

Distribution List for ARPA Order 2840 Unclassified Contracts

Raytheon Co. 28 Seyon Street Waltham, MA 02164 Attn: Dr. H. Statz	1 copy
Raytheon Co. Bedford Labs Missile Systems Division Bedford, MA 01730 Attn: Dr. H A. Mehlhorn	1 copy
Riverside Research Institute 80 West End Street New York, N. Y. 10023 Attn: Dr. L. H. O'Neill (HPEGL Library)	1 copy 1 copy
R&D Associates, Inc. P.O. Box 3580 Santa Monica, CA 90431 Attn: Dr. R. E. Lelevier Dr. R. Hundley	1 copy 1 copy
Rockwell International Corp. Rocketdyne Division Albuquerque District Office 3636 Menaul Blvd., NE, Suite 211 Albuquerque, N.M. 87110 Attn: C. K. Kraus, Mgr.	1 copy
SANDIA Corp. P.O. Box 5800 Albuquerque, N.M. 87115 Attn: Dr. Al. Narath	1 copy
Stanford Research Institute Menlo Park, CA 94025 Attn: Dr. R. A. Armistead Dr. J. E. Malick	1 copy 1 copy
Science Applications Inc. 1911 N. Ft. Myer Drive Arlington, VA 22209 Attn: L. Peckham	1 copy
Science Applications Inc. P.O. Box 328 Ann Arbor, MI 48103 Attn: R. E. Meredith	1 copy

Distribution List for ARPA Order 2840 Unclassified Contracts

Science Applications Inc.
3 Preston Court
Bedford, MA 01703
Attn: Dr. F. A. Horrigan

1 copy

Science Applications Inc.
P.O. Box 2351
La Jolla, CA 92037
Attn: Dr. John Asmus

1 copy

Systems, Science and Software
P.O. Box 1620
La Jolla, CA 92037
Attn: Alan F. Klein

1 copy

Systems Consultants, Inc.
1050 31st Street, N.W.
Washington, D. C. 20007
Attn: Dr. R. B. Keller

1 copy

Thiokol Chemical Corp.
WASATCH Division
P.O. Box 524
Brigham City, UT 84302
Attn: Mr. J. E. Hansen

1 copy

TRW Systems Group
One Space Park
Bldg. R-1, Rm. 1050
Redondo Beach, CA 90278
Attn: Mr. Norman Campbell

1 copy

United Aircraft Research Labs
400 Main Street
East Hartford, CT 06108
Attn: Mr. G. H. McLafferty
Mr. Albert Angelbeck

3 copies
1 copy

United Aircraft Corp.
Pratt and Whitney Acft. Div.
Florida R&D Center
West Palm Beach, FL 33402
Attn: Dr. R. A. Schmidtke
Mr. Ed. Pinsley

1 copy
2 copies

VARIAN Associates
EIMAC Division
301 Industrial Highway
San Carlos, CA 94070
Attn: Mr. Jack Quinn

1 copy

Distribution List for ARPA Order 2840 Unclassified Contracts

Vought Systems Division LTV Aerospace Corp. P.O. Box 5907 Dallas, TX 75222 Attn: Mr. F. G. Simpson Mail Station 254142	1 copy
Westinghouse Electric Corp. Defense and Space Center Friendship International Airport - Box 746 Baltimore, MD 21203 Attn: Mr. W. F. List	3 copies
Westinghouse Research Labs Beulah Road, Churchill Boro Pittsburgh, PA 15235 Attn: Dr. E. P. Riedel Mr. R. L. Hundstad	1 copy 1 copy
United Technologies Research Center East Hartford, CT 06108 Attn: A. J. DeMaria W. Glenn	1 copy 1 copy
Stanford Research Institute Menlo Park, CA 94025 Attn: Dr. F. T. Smith	1 copy
Cleveland State University Cleveland, OH 44115 Attn: Dean Jack Soules	1 copy
Exxon Research and Engineering Co. P.O. Box 8 Linden, N. J. 07036 Attn: D. Grafstein	1 copy
University of Maryland Dept. of Physics and Astronomy Attn: D. Currie College Park, MD 20742	1 copy
North American Rockwell Corp. Autonetics Division Attn: R. Gudmundsen 3370 Miraloma Avenue Anaheim, CA 92803	1 copy

Distribution List for ARPA Order 2840 Unclassified Contracts

Massachusetts Institute of Technology 77 Massachusetts Avenue Cambridge, MA 02138 Attn: Prof. A. Javan Prof. C. Dewey Prof. B. Lax	1 copy 1 copy 1 copy
Lockheed Missile and Space Co. Palo Alto Research Laboratories Attn: Dr. R. C. Ohlman Palo Alto, CA 94304	1 copy
University of Texas at Dallas Attn: Prof. Carl B. Collins P.O. Box 30365 Dallas, TX 75230	1 copy
Stanford University Microwave Laboratory Attn: S. E. Harris Stanford, CA 94305	1 copy
Northrop Corporation 3401 West Broadway Hawthorne, CA 90250 Attn: M. Bhau mik M. Mann E. Ault	1 copy 1 copy 1 copy
Polytechnic Institute of New York Aerophysics Institute Route 110 Farmingdale, NY 11735 Attn: M. Bloom W. T. Walter	1 copy 1 copy
Mathematical Sciences Northwest, Inc. Attn: Dr. S. R. Byron 4545 Fifteenth Ave., N. E. Seattle, Washington 98105	1 copy
Wayne State University Research Institute for Engineering Sciences Attn: Dr. E. R. Fisher Detroit, MI 48202	1 copy
University of Washington College of Engineering Attn: W. H. Christiansen Seattle, WA 98195	

Distribution List for ARPA 2940 Unclassified Contracts

Univ. of Southern California Electronic Sciences Lab University Park Los Angeles, CA 90007 Attn: Dr. Curt Whittig Dr. S. R. Leone	1 copy 1 copy
Yale University Dept. of Engineering and Applied Sciences Attn: Dr. A. Herzenberg New Haven, Conn. 06520	1 copy
Columbia University Dept. of Chemistry 405 Low Memorial Library New York, N.Y. 10027 Attn: Dr. G. Flynn Dr. R. N. Zare	1 copy 1 copy
Cutler-Hammer, Inc. AIL Division Attn: Dr. Alex Stein Deer Park, L.I., N.Y. 11729	1 copy
Kansas State University Dept. of Chemistry Attn: Dr. D. Setser Manhattan, Kansas 66506	1 copy
Vanderbilt University Dept. of Chemistry Attn: Dr. J B. Tellinghuisen Nashville, Tenn. 37235	1 copy
Maxwell Laboratories Attn: Mr. Robert Hunter 9244 Balboa Avenue San Diego, CA 92123	1 copy
New York Area Office Office of Naval Research, Scientific Dept. 715 Broadway (5th Floor) Attn: Dr. Irving Rowe New York, N. Y. 10003	1 copy